

This is a repository copy of *Synthesis, chemistry, physicochemical properties and industrial applications of amino acid surfactants : A review.*

White Rose Research Online URL for this paper:

<https://eprints.whiterose.ac.uk/133786/>

Version: Accepted Version

Article:

Tripathy, Divya, Mishra, Anuradha, Clark, James Hanley orcid.org/0000-0002-5860-2480 et al. (1 more author) (2018) *Synthesis, chemistry, physicochemical properties and industrial applications of amino acid surfactants : A review.* *Comptes Rendus Chimie.* pp. 112-130. ISSN 1878-1543

<https://doi.org/10.1016/j.crci.2017.11.005>

Reuse

This article is distributed under the terms of the Creative Commons Attribution-NonCommercial-NoDerivs (CC BY-NC-ND) licence. This licence only allows you to download this work and share it with others as long as you credit the authors, but you can't change the article in any way or use it commercially. More information and the full terms of the licence here: <https://creativecommons.org/licenses/>

Takedown

If you consider content in White Rose Research Online to be in breach of UK law, please notify us by emailing eprints@whiterose.ac.uk including the URL of the record and the reason for the withdrawal request.

Manuscript Number: CRCHIMIE-D-17-00356R1

Title: Synthesis, Chemistry, Physicochemical Properties and Industrial Applications of Amino Acid Surfactants: A Review

Article Type: Revue / Account

Keywords: Aminoacid surfactants, protein surfactants, microbial surfactants, gemini surfactants, bolaamphiphiles, biodegradable surfactants.

Corresponding Author: Professor Anuradha Mishra, PhD

Corresponding Author's Institution: School of Vocational Studies & Applied Sciences, Gautam Buddha University

First Author: Anuradha Mishra, PhD

Order of Authors: Anuradha Mishra, PhD; Divya B Tripathy, Ph.D; James Clark, Ph.D; Thomas Farmer, Ph.D

Manuscript Region of Origin: AFGHANISTAN

Abstract: Surfactant use throughout mankind is extensive, from their initial applications as detergents extending into use in medicine, lubricant, cosmetics and even enhanced oil recovery. However, the image of surfactant use has in the past been tarnished by of issues with low biodegradability and their synthesis from non-sustainable resources. Amino acid-based surfactants are a class of surfactants derived from a hydrophobe source coupled with simple amino acids, mixed amino acids from synthesis or from protein hydrolysates, and as such can be derived solely from renewable resources. There are several pathways for their synthesis and this allows for extensive structural diversity in this class of surfactants, resulting in widespread tuneable functionality in their physiochemical properties. This review includes the details of most of the available routes of synthesis for amino acid surfactants and the impact of the diverse routes on their final physiochemical properties, including solubility, dispersability, toxicity and biodegradability. The diversity offered by the structural variation in amino acids surfactants offers many exciting commercial opportunities for this ever growing class of surfactants. It also includes a discussion on current and future potential uses of amino acid surfactants.

Suggested Reviewers:

Opposed Reviewers:

**Synthesis, Chemistry, Physicochemical Properties and Industrial Applications of Amino acid
Surfactants: A Review**

Divya Bajpai Tripathy¹, Anuradha Mishra^{1*}, James Clark² and Thomas Farmer²

¹Department of Applied Chemistry, School of Vocational Studies and Applied Sciences,
Gautam Buddha University, Greater Noida-201312, India

²Green Chemistry Centre of Excellence, Department of Chemistry, University of York,
Heslington, York, YO10 5DD, UK

Abstract

Surfactant use throughout mankind is extensive, from their initial applications as detergents extending into use in medicine, lubricant, cosmetics and even enhanced oil recovery. However, the image of surfactant use has in the past been tarnished by of issues with low biodegradability and their synthesis from non-sustainable resources. Amino acid-based surfactants are a class of surfactants derived from a hydrophobe source coupled with simple amino acids, mixed amino acids from synthesis or from protein hydrolysates, and as such can be derived solely from renewable resources. There are several pathways for their synthesis and this allows for extensive structural diversity in this class of surfactants, resulting in widespread tuneable functionality in their physiochemical properties. This review includes the details of most of the available routes of synthesis for amino acid surfactants and the impact of the diverse routes on their final physiochemical properties, including solubility, dispersability, toxicity and biodegradability. The diversity offered by the structural variation in amino acids surfactants offers many exciting commercial

opportunities for this ever growing class of surfactants. It also includes a discussion on current and future potential uses of amino acid surfactants.

Keywords: Amino acid surfactants, protein surfactants, microbial surfactants, gemini surfactants, bolaamphiphiles, biodegradable surfactants.

*Corresponding author E-mail: anuradha_mishra@rediffmail.com

Table of content

1. Introduction

2. History

3. Structural properties

4. Chemistry

5. Classification

5.1. On the basis of origin

5.1.1. Natural

5.1.2. Synthetic

5.2. On the basis of the substitution of aliphatic chain

5.2.1. On the basis of substitution site

5.2.1.1. N-substituted compounds

5.2.1.2. C-substituted compounds

5.2.1.3. N- and C- substituted compounds

5.3. On the basis of number of hydrophobic tail

5.4. On the basis of head group type

1
2
3
4 5.4.1.1. Anionics
5
6

7 5.4.1.2. Cationics
8

9 5.4.1.3. Amphoterics
10

11 5.4.1.4. Nonionics
12
13

14 **6. Synthesis**

15
16 6.1.1. Fundamental synthesis paths
17

18
19 6.1.2. Development in the synthesis/production
20

21 6.1.2.1. Synthesis of single-chain amino acid/peptide surfactants
22

23 6.1.2.2. Synthesis of gemini amino acid/peptide surfactants
24

25 6.1.2.3. Synthesis of glycerolipid amino acid/peptide surfactants
26

27 6.1.2.4. Synthesis of bolaamphiphiles amino acid/peptide surfactants
28
29
30

31 **7. Physiochemical properties**

32
33 7.1.1. Solubility
34

35
36 7.1.2. Thermal stability
37

38 7.1.3. Critical micelle concentration (CMC)
39

40
41 7.1.4. Krafft temperature and Krafft point
42

43 7.1.5. Surface tension
44

45 7.1.6. Binary surface interaction
46

47 7.1.7. Aggregation properties
48

49 7.1.8. Interfacial adsorption
50

51 7.1.9. Phase behaviour
52

53 7.1.10. Emulsification
54

55 7.1.11. Dispersability
56
57
58
59
60
61
62
63
64
65

8. Toxicity

9. Antimicrobial activity

10. Biodegradability

11. Haemolytic activity

12. Rheological Properties

13. Industrial applications

13.1. In agriculture

13.2. In commercial laundry detergents

13.3. In lubricants

13.4. In medicine

13.4.1. Drug carriers and preparation of functional liposomes

13.4.2. Gene therapy and DNA transfection

13.4.3. Antiviral agents

13.5. In food processing industry

13.6. In cosmetic industry

13.7. In microbial enhanced oil recovery

13.8. In nanomaterials

13.9. Other applications

14. Current challenges

15. Conclusions

1. Introduction

Surfactants are the group of organic compounds which continue to attain great interest from researchers due to their wide range of applications including as laundry

1
2
3
4 detergents, emulsifiers, corrosion inhibitors, oil recovery and pharmaceuticals. These are the
5
6 most representative chemical products to be consumed in major quantities daily and globally
7
8 and have in the past led to adverse effects on the aquatic environment. Many studies have
9
10 previously revealed the adverse impact of widespread use of conventional surface-active
11
12 agents on the environment (Ivanković & Hrenović, 2010). As such non-toxicity,
13
14 biodegradability and biocompatibility of surfactants have become almost equally important
15
16 for the consumers as their functional performance.
17
18
19
20

21 Biosurfactants are a class of green and sustainable surface active agents naturally
22
23 synthesized from microorganisms such as bacteria, fungi and yeast or excreted
24
25 extracellularly. Synthetic equivalents to biosurfactants can therefore be prepared by
26
27 designing molecules that imitate natural amphiphilic structures such as phospholipids,
28
29 alkylglucosides and acyl amino acids. Amino acid surfactants (AAS) are one such type of
30
31 surfactant which can generally be originated from animal or agricultural-derived feedstocks.
32
33 AAS have been gaining great interest of scientists over the last two decades as novel
34
35 surfactants as they can be synthesized using renewable sources and their ease of
36
37 degradability and harmless by-product make them safer for our environment (Kango, 2010).
38
39
40
41
42

43 AAS can be defined as a group of surfactants made up of amino acids comprising of
44
45 amino acid group ($\text{HO}_2\text{C}-\text{CHR}-\text{NH}_2$) or its residue ($\text{HO}_2\text{C}-\text{CHR}-\text{NH}_2$). These two functional
46
47 regions of amino acids give the possibility to derive an extensive range of surfactants. There
48
49 are total of 20 standard proteinogenic amino acids known in nature which are responsible
50
51 for growth and all physiological reactions of living kingdom. They differ from each other
52
53 only on the basis of the residue, R (Taubeneck, 1977) (Fig .1). Some are non-polar and
54
55 hydrophobic, others are polar and hydrophilic, some are basic and some are acidic. As
56
57
58
59
60
61
62
63
64
65

1
2
3
4 amino acids are renewable compounds, surfactants synthesized from amino acids also have
5
6 great potential as sustainable and eco-friendly substances (Infante *et al.*, 1997). Simple and
7
8 natural structure, low toxicity and fast biodegradation often make them superior over their
9
10 conventional counterparts. Their production can be via different biotechnological and
11
12 chemical routes using renewable raw materials such as amino acids and vegetable oils.
13
14

15
16 Amino acids were first discovered as a substrate for surfactants in the early 20th
17
18 century (Morán *et al.*, 2004). Primarily they were used as preservatives in pharmaceuticals
19
20 and cosmetic formulations. Furthermore, they were found to be biologically active against a
21
22 variety of disease-causing bacteria, tumors, and viruses (Husmann, 2008). In 1988
23
24 availability of amino acid surfactants at low cost increased the interest of researchers to
25
26 study their surface activity (Takehara, 1989). Nowadays, along with the growth of
27
28 biotechnology, a few amino acids are also able to be synthesized commercially at large scale
29
30 by yeasts, thus justifying their production to be more environmental friendly (Diniz Rufino
31
32 *et al.*, 2014).
33
34
35
36
37

38 **2. History**

39

40
41 The value of the structures of naturally occurring amino acids as raw materials for
42
43 preparing amphiphiles was predicted as soon as they were discovered early in the 19th
44
45 century. The first research on amino acid surfactant synthesis was reported in 1909 by
46
47 Bondi (Bondi, 1909). In this research N-acylglycine and N-acylalamine as hydrophilic
48
49 moieties of surfactants have been introduced. Subsequent work involved the synthesis of
50
51 lipo-amino acid using glycine and alanine (Gallot & Hassan, 1989). Hentrich and co-workers
52
53 published a series of findings and filed the first patent on the applications of acylsarcosinate
54
55 and acylaspartate as surfactants in household cleaning products such as shampoos,
56
57
58
59
60
61
62
63
64
65

detergents and toothpastes (Hentrich *et al.*, 1936). Subsequently the synthesis and physicochemical properties of acyl amino acids have been studied by many researchers (Heitmanr, 1968; Presenz, 1996). To date a significant number of research publications have been published on the synthesis (Heitmanr, 1968; Perez *et al.*, 2009; Lundberg *et al.*, 2011; Seguer *et al.*, 1994a; Seguer *et al.*, 1994b; Piera *et al.*, 1998; Clapes & Infante, 2002), properties (Franklin & Snow 1981; Solans *et al.*, 1989; Kunieda *et al.*, 1992; Fördedal *et al.*, 1993; Wang *et al.*, 2001; Dana *et al.*, 2011), industrial applications (Infante *et al.*, 1984; Infante *et al.*, 1992; Infante *et al.*, 1994) and biodegradability (Piera *et al.*, 1998; Clapes *et al.*, 1999; Burczyk, 2003; Trivedi *et al.*, 2011; Pinazo, 2011) of amino acid surfactants.

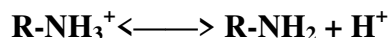
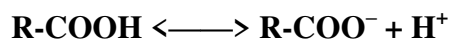
3. Structural properties

Non-polar hydrophobic fatty acid chains of amino acid surfactants may vary in their structure, length and number. Structural variety and high surface activities of amino acid surfactants explain their wide compositional diversity and broad range of physicochemical and biological properties. The difference in the amino acid or peptide head group of amino acid surfactants determines the adsorption, aggregation and biological activity of these surfactants. Their types in terms of cationic, anionic, non-ionic and amphoteric, depends on the functional groups present in them. The combination of hydrophilic polar amino acids and non-polar long chain hydrophobic moiety for building up the amphiphilic structure has created molecules with high surface activity (Infante *et al.*, 1997; Rentsch, 2002; Abdullah *et al.*, 2009). In addition asymmetric carbons present in the molecule help in the formation of chiral molecules (Solans *et al.*, 1990; Walther & Netscher, 1996).

4. Chemistry

All peptides and polypeptides are the polymerization product of 20 or so proteinogenic α -amino acids. All 20 α -amino acids are comprised of a carboxylic acid ($-\text{COOH}$) functional group and an amino ($-\text{NH}_2$) functional group attached to the same tetrahedral α -carbon atom (Morán *et al.*, 2004). Amino acids differ from each other on the basis of distinct R-groups attached to the α -carbon (except in the case of glycine where the R-group is hydrogen). R groups may differ from each other in structure, size and electric charge (acidity/basicity). These differences also determine the solubility of amino acids in water.

Presence of 4 different substituent attached to α -carbon makes amino acids chiral in nature (except glycine) and hence optically active. Amino acids have two possible configurations; both are non-superimposable mirror images of each other, though in nature the L-stereoisomers are significantly more abundant. Presence of aromatic R-groups in some amino acids (phenylalanine, tyrosine and tryptophan) results in the absorption of ultraviolet light with an absorbance maximum in the range of 280 nm (Solans *et al.*, 1990). Both, the acidic α -COOH and basic α -NH₂ groups present in amino acids are able to ionize and create ionic equilibrium as follows:



Amino acids contain at least two weakly acidic groups; however, the carboxyl group is a far stronger acid than the amino group. At pH 7.4, the carboxyl group is unprotonated and the amino group protonated. An amino acid with no ionizable R-group would be electrically neutral at this pH and form zwitterions (Tamarkin *et al.*, 2014).

5. Classification

Four parameters are taken into consideration to classify amino acid surfactants, these are sequentially addressed below:

5.1 On the basis of origin

On the basis of their origin amino acid surfactants can be categorized into two classes:

5.1.1 Natural (Pathak & Keharia, 2014; Walia & Cameotra, 2015)

Some naturally occurring amino acid containing compounds also possess the ability to reduce surface and interfacial tension, some even surpassing the efficacy of glycolipids. These types of amino acid surfactants are also known as lipopeptides. Lipopeptides are low molecular weight compounds, which are generally produced by *Bacillus* species. This class of amino acid surfactants are further classified into three subclasses: surfactin, iturin and fengycin (Fig. 2).

The surfactin (Hathout *et al.*, 2000; Arima *et al.*, 1968) family covers the heptapeptide variants of the esperin, pumilacidin, lichenysin and surfactin groups. In such types of surfactants, the peptide moiety is linked with a C₁₂-C₁₆ unsaturated linear, iso or anteiso β -hydroxyl fatty acid chain. Surfactins are the macrolactone rings, in which ring closure is catalysed in between the β -hydroxyl fatty acid and the C-terminal peptide.

In the iturin (Peypoux *et al.*, 1999) subclass, there are mainly six variants i.e. iturin A and C, mycosubtilin and bacillomycin D, F and L. In all the cases, the heptapeptide is attached with a varied C₁₄-C₁₇ chain of β -amino fatty acids. In Iturins, because of its β -amino nature, an amide bond is formed with the C-terminal group and thus forms a macrolactame structure.

The Fengycins (Winkelmann *et al.*, 1983; Ongena *et al.*, 2007; Kim *et al.*, 2010) subclass encompasses fengycins A and B, which are also known as plipastatins when Tyr⁹

1
2
3
4 is D-configured. These decapeptides are also linked with a C₁₄-C₁₈, saturated or unsaturated
5
6 linear, iso or anteiso-β-hydroxyl fatty acid chain. Structurally, fengycins are also a
7
8 macrolactone ring that includes a Tyr side chain at third position in the peptide sequence
9
10 and form an ester bond with the C-terminal residue, thus form an internal ring like structure
11
12 in many *Pseudomonas* lipopeptides.
13
14

15 16 **5.1.2 Synthetic (Xia, 2001)** 17

18
19 Amino acid surfactants can also be synthesized experimentally by use of any of the
20
21 acidic, basic and neutral amino acids. Common amino acids used for their synthesis are
22
23 glutamic acid, serine, proline, aspartic acid, glucine, arginine, alanine, leucine and protein
24
25 hydrosylates. Surfactants of this subclass can be prepared chemically, enzymatically and
26
27 chemoenzymatically however, chemical synthesis has been found more economically
28
29 feasible for their production. Common examples of this include N-lauroyl-L-glutamic acid,
30
31 N-palmitoyl-L-glutamic acid.
32
33

34 35 **5.2. On the basis of the substitution of aliphatic chain** 36

37
38 On the basis of the substitution of aliphatic chain amino acid based surfactants can
39
40 be classified into two types (Clapes & Infante, 2002):
41
42

43 44 **5.2.1. On the basis of substitution site** 45

46 In this type, there are three different categories:
47

48 49 **5.2.1.1. N-substituted amino acid surfactants** 50

51 In N-substituted compounds an amino group is substituted either with a lipophilic
52
53 moiety or a carboxylic group, resulting in the loss of alkalinity. The simplest examples of N-
54
55 substituted amino acid surfactants are N-acyl amino acids, which are anionic in nature. N-
56
57 substituted amino acid surfactants have an amide linkage formed in between the
58
59
60
61
62
63
64
65

hydrophobic moiety and the hydrophilic moiety. The ability of amide linkage to form hydrogen bond eases the degradation of such surfactants under acidic environment and hence makes them biodegradable.

5.2.1.2. C-substituted amino acid surfactants

In C-substituted compounds substitution takes place on the carboxylic group through an amide or an ester bond. C-substituted compounds are typically cationic in nature such as esters or amides.

5.2.1.3. N- and C- substituted amino acid surfactants

In second type of the surfactants, both the amino and the carboxylic groups are present as part of the hydrophilic moiety. These types are amphoteric in nature.

5.3 On the basis of number of hydrophobic tail

On the basis of number of hydrophobic tail and head groups, amino acid surfactants are classified into four groups (Moran *et al.*, 2004). Single or linear chain amino acid surfactants, dimeric or gemini amino acid surfactants, glycerolipids type amino acid surfactants and bolaamphiphiles amino acid surfactants. Linear chain surfactants are the surfactants, comprised of amino acids with the minimum one hydrophobic tail (Fig. 3). Dimeric amino acid surfactants have two amino acids polar head group and two hydrophobic tails per molecule (Fig. 4). In this class, two linear amino acids surfactants are joined together via spacer and hence also called dimerics. Glycerolipids type amino acid surfactants contain two hydrophobic tails attached to a one common amino acid head group (Fig.5). These surfactants can be considered as an analogue of monoglycerides, diglycerides and phospholipids. Bolaamphiphiles amino acid surfactants are a group of surfactants that has two polar amino acid heads that are joined together via single hydrophobic tails (Fig. 6).

5.4. On the basis of head group type

On the basis of polar head group amino acid based surfactants are further classified into cationic, anionic, zwitterionic and non-ionic amino acid surfactants.

5.4.1 Cationic

In this class of surfactants, the head group bears positive charge. The first commercially available cationic amino acid surfactant is cocoyl arginine ethyl ester (CAE) as a PCA (pyrrolidone carboxylic acid) salt. This surfactant has unique and varied properties that make it useful in disinfectants, antimicrobial agents, antistatic agents, hair conditioner and it is also very mild to eyes and skin and easily biodegradable (Xia., 2001). Singare & Mhatre, 2012 have synthesized arginine based cationic amino acid surfactants and evaluated them on the basis of their physicochemical properties. In this study, they claimed the good yield of product using Schotten Baumann reaction conditions. Increase in the surface activity and decrease in the critical micelle concentration (CMC) of surfactants was found with the increase in alkyl chain length and the hydrophobicity. Quaternary acyl protein is another known example of this class which is generally used as conditioner in hair care products (Kim *et al.*, 2010).

5.4.2. Anionic

In anionic surfactants, negative charge is present on the polar head group of surfactants. Sarcosinate surfactants are an example of anionic amino acid surfactants. Sarcosine ($\text{CH}_3\text{-NH-CH}_2\text{-COOH}$, *N*-methyl glycine) is an amino acid normally found in sea urchins and starfish and is chemically related to glycine ($\text{NH}_2\text{-CH}_2\text{-COOH}$), a basic amino acid found in the cells of mammals. Lauric acid, myristic acid and oleic acid and their halides and esters are generally used to synthesize sarcosinate surfactants. Sarcosinate are

1
2
3
4 mild in nature, due to which they are generally used in mouthwashes, shampoos, aerosol
5
6 shaving lathers, sunscreens, skin cleansers and other cosmetics (Castillo *et al.*, 2001).
7
8

9 Other commercial anionic amino acid surfactants include Amisoft CS-22 and
10
11 Amilite GCK-12, which are the trade names of Sodium N-Cocoyl-L-Glutamate and
12
13 Potassium N-Cocoylglycinate, respectively. Amilite is generally used as foaming agent,
14
15 detergent, solubilizer, emulsifier and as a dispersing agent and has numerous applications in
16
17 cosmetics as shampoo and bath soap, body wash, toothpaste, face wash, facial soap, contact
18
19 lens cleaners and household tensides (Aubert & Dussault, 2008). Amisoft is used as a mild
20
21 skin and hair cleansers and predominantly used in face and body cleansers, synthetic
22
23 detergent bars, body care hair shampoos and other skin care products (Castillo *et al.*, 2000).
24
25
26
27

28 **5.4.3. Zwitterionic (Amphoteric)**

29
30

31 Amphoteric surfactants contain both acidic and basic sites and thus change charge by
32
33 varying pH. They behave like anionics in alkaline medium, behave as cationics in acidic
34
35 environment and amphoteric in neutral medium. Lauroyl lysine (LL) and Alkoxy (2-
36
37 hydroxypropyl) arginine are the only known amino acid based amphoteric surfactants. LL is a
38
39 condensation product of lysine and lauric acid. Due to its amphoteric structure, LL is
40
41 insoluble in almost all type of solvents except highly alkaline or acidic solvents. Excellent
42
43 adhesive property towards hydrophilic surfaces and low friction coefficient of LL as organic
44
45 powder imparts excellent lubricating ability in this surfactant. LL is widely used in skin
46
47 creams and hair conditioners or as a lubricant (Mohr *et al.*, 2000).
48
49
50
51
52

53 **5.4.4. Nonionic**

54

55 Nonionic surfactants are characterized with polar head groups that carry no formal
56
57 charge. Sabagh and his co-workers prepared eight novel ethoxylated nonionic surfactants
58
59
60
61
62
63
64
65

1
2
3
4 using oil soluble α -amino acids. In this process, esterification of L-phenylalanine and L-
5
6 leucine followed by amidation with cetyl alcohol and palmitic acid, respectively yielded
7
8 two amides and two esters of α -amino acids. The ethylene oxide was then condensed with
9
10 the amides and esters prepared and produced three different polyethylene oxide units of 40,
11
12 60, and 100 as phenylalanine derivatives. These nonionic type amino acid surfactants were
13
14 found to have good detergency and foaming properties (Sabagh *et al.*, 2009).
15
16
17

18 **6. Synthesis**

19 **6.1 Fundamental synthesis paths**

20
21 In amino acid surfactants, the hydrophobic group may be attached either at the amine
22
23 moiety, the carboxylic acid moiety, or through the side chain of the amino acid. On the basis
24
25 of this approach, four fundamental synthesis paths are available to researchers (Kimura *et*
26
27 *al.*, 1992) as given in figure 7.
28
29
30
31

32
33 **Path 1** shows the production of amphiphilic esteramine through esterification
34
35 reactions, in which surfactants synthesis is typically achieved by refluxing fatty alcohols and
36
37 amino acid in the presence of dehydrating agent and acidic catalyst. In some reactions
38
39 sulphuric acid plays the role of both catalyst and dehydrating agent.
40
41
42

43 **Path 2** shows the synthesis of amphiphilic amidoamine by creating the amide bond
44
45 through the reaction of an activated amino acid and alkyl amine to create the desired amide
46
47 bond.
48
49

50 **Path 3** shows the production of amidoacid through the reaction of amine group of
51
52 amino acid with fatty acid.
53
54

55 **Path 4** shows the synthesis of long chain alkyl amino acids via the reaction of amine
56
57 group with alkyl halogen.
58
59
60
61
62
63
64
65

1
2
3
4 **Path 5** involves the coupling of the specific function of the side group of the amino
5
6 acid. In this type of path, generally carboxylic group of aspartic acid and glutamic acid react
7
8 with fatty alcohol to form anhydride (Fig.8).
9

10 11 **6.2. Development in the synthesis/production**

12 13 **6.2.1. Synthesis of single-chain amino acid/peptide surfactants**

14
15 Synthesis of N-Acyl and O-acyl amino acids (Kimura *et al.*, 1992) or peptides can be
16
17 achieved by the enzyme-catalysed acylation reaction of amine or alcohol groups with fatty
18
19 acids. The first description of solvent-free lipase catalyzed synthesis of amino acid amide or
20
21 methyl ester derivatives was reported by using *Candida antarctica* with the yield in the
22
23 range of 25-90% depending on the target amino acid (Godtfredsen and Bjoerkling, 1990). In
24
25 some reactions, ethylmethyl ketone has also been used as solvent. Vonderhagen *et al.*, 1999
26
27 also illustrated the lipase and proteases catalyzed N-acylation of amino acids, protein
28
29 hydrolysates and/or their derivatives using the mixture of water with organic solvents such
30
31 as dimethylformamide/water and butylmethylketone. In these earlier examples, the major
32
33 issue with enzymatic synthesis of amino acid surfactants was typically very low yields.
34
35 Valivety *et al.*, 1997 reported yields of just 2– 10% of N-myristoyl amino acid derivatives
36
37 even after many days under incubation at 70°C by using different lipases. Problems
38
39 associated with the low yields of amino acid were also supported by Montet and his co-
40
41 workers (Montet *et al.*, 1990) during the synthesis of N-acyllysines using fatty acids and
42
43 vegetable oils. They reported a maximum 19% yield of the product under both the solvent-
44
45 free conditions as well as under organic solvents. Valivety *et al.*, 1998 also supported the
46
47 same issue while synthesizing N-Cbz-L-Lysine or N-Cbz- Lysine methyl ester derivatives.
48
49 In this study, they claimed 80 % yield of 3-O-myristoyl-L-serine using N-protected serine as
50
51
52
53
54
55
56
57
58
59
60
61
62
63
64
65

1
2
3
4 substrate and Novozyme 435 as catalyst under melted solvent free environment. Nagao and
5
6 Kito studied the results of O-acylation of L-serine, L-homoserine, L-threonine and L-
7
8 tyrosine using lipases obtained by *Candida cylindracea* and *Rhizopusdelemar* in aqueous
9
10 buffered medium and reported the acylation of L-homoserine and L-serine upto some extent
11
12 with low yield whereas no acylation takes place on L-threonine and L-tyrosine (Nagao &
13
14 Kito, 1989). Various researchers supported the use of cheap and easily available substrate to
15
16 synthesize cost effective amino acid surfactants. Soo'sresearch group claimed the best
17
18 results of palm oil based surfactants using immobilized lipoenzyme. In this investigation
19
20 they stated that although reactions with palm oil fractions required long reaction duration of
21
22 6 days but resulted better yields of products (Soo *et al.*, 2003). Gerova and his co-workers
23
24 studied the synthesis and surface active properties of several optically active and racemic
25
26 mixtures chiral of *N*-palmitoyl amino acid surfactants based on methionine, proline, leucine,
27
28 threonine, phenylalanine and phenylglycine (Gerova *et al.*, 2008). Pang & Chu, 2010
29
30 described theco-polycondensation of amino acid based monomers and dicarboxylic acid
31
32 based monomers to synthesize a series of biodegradable functional amino acid-based
33
34 polyesteramides in solution.
35
36
37
38
39
40
41
42

43 Cantacuzene and Guerreiro reported the esterification of carboxylic acid group of
44
45 Boc-Ala-OH and Boc-Asp-OH using long chain aliphatic alcohols and diols using
46
47 dichloromethane and Sepharose 4B as solvent and catalyst, respectively. In this research,
48
49 good yields (51%)of Boc-Ala-OH was obtained with fatty alcohols up to 16 carbon atoms
50
51 whereas 6 and 12 carbon atoms were found better and gave 63% of Boc-Asp-
52
53 OH(Cantacuzene & Guerreiro, 1989).Clapes *et al.* claimed 58–76% yield and 99.9% purity
54
55 of N-arginine alkyl amide derivatives obtained through the formation of amide and ester
56
57
58
59
60
61
62
63
64
65

1
2
3
4 bonds between Cbz-Arg-OMe and various long chain alkyl amines and fatty alcohols in the
5
6 presence of Papain from *Carica papaya* latex as catalyst (Clapes *et al.*, 1999).
7
8

9 **6.2.2. Synthesis of gemini amino acid/peptide surfactants**

10
11 Amino acid based gemini surfactants are comprised of two linear amino acid
12
13 surfactants molecules joined together head to head via a spacer (Yoshimura *et al.*, 2012).
14
15 Two possible Schemes have been derived for the chemo-enzymatic synthesis of gemini
16
17 types amino acid based surfactants (Schemes1, 2). In the first Scheme, two amino acid
18
19 derivatives react with the spacer followed by the introduction of the two hydrophobic
20
21 groups. In the second Scheme, the two linear structures are joined directly through
22
23 bifunctional spacer chain (Piera *et al.*, 2000). Valivety and his co-workers (Valivety *et al.*,
24
25 1997) were the first to develop the enzymatic synthesis of gemini lipoamino acids.
26
27 Yoshimura and his co-worker (Yoshimura *et al.*, 2012) studied the synthesis, adsorption,
28
29 and aggregation properties of an amino acid-based gemini surfactant based on cystine and n-
30
31 alkyl bromide. Comparisons of the synthesized surfactants have also been performed against
32
33 their corresponding monomeric counterparts. Faustino *et al.*, 2009 described the synthesis
34
35 of anionic urea-based monomeric amino acid surfactants and their corresponding geminis
36
37 based on L-cystine, D-cystine and DL cystine, as well as derived from L-cysteine, L-
38
39 methionine and L-cysteic acid and their characterization was achieved in terms of electrical
40
41 conductivity, equilibrium surface tension, and steady-state fluorescence spectroscopy
42
43 techniques. Comparison studies between monomeric forms and geminis stated that geminis
44
45 have lower critical micelle concentration (CMC) values.
46
47
48
49
50
51
52
53
54

55 **6.2.3. Synthesis of glycerolipidamino acid/peptide surfactants**

56
57
58
59
60
61
62
63
64
65

Glycerolipidamino acid/peptide surfactants comprises of a novel class of lipoamino acids, which are the structural analogues of mono-diacylglycerides and phospholipids as one or two aliphatic chains and one amino acid joined together via ester bonds in the glycerol backbone. Synthesis of such types of surfactants starts with the preparation of glyceryl esters of amino acid in the presence of an acid catalyst such as BF_3 at elevated temperatures (Valivety *et al.*, 1997). Enzymatic synthesis using hydrolases, proteases and lipases as catalyst was also found to be good alternative (Scheme 3) (Mitin *et al.*, 1997). Enzymatic synthesis of dilauroylated arginine glyceride conjugate with papain enzyme has also been reported. Synthesis of diacylglycerolipid conjugates from acetyl-arginine and evaluation of their physicochemical properties have also been reported (Moran *et al.*, 2001; Perez *et al.*, 2004).

6.2.4. Synthesis of bolaamphiphiles amino acid/peptide surfactants

Amino acid-based bolaamphiphiles contain two amino acids connected with a hydrophobic linker. Franceschi *et al.*, 1999 described the synthesis of bolaamphiphiles with two amino acids (D- or L-alanine or L-histidine) and an alkyl chain of varying length and studied their surface active properties. In this research work, they discussed the synthesis and aggregation of novel bolaamphiphiles bearing amino acid moieties using unusual β -amino acids or an alcohol and spacers with C_{12} - C_{20} . The unusual β -amino acids used may be sugar amino acid, an AZT-derived amino acid, a norbornene amino acid, and an AZT-derived amino alcohol (Scheme 4). Polidori *et al.*, 2006 has also given synthesis of symmetric bolaamphiphiles derived from tris (hydroxymethyl) aminomethane (Tris) (Scheme 4).

7. Physiochemical Properties

Amino acid based surfactants are known to have wide range of desirable properties that increase their applicability in varied applications such as good solubilization, good emulsifying properties, high efficiency, high surface activity and good lime resistance (calcium tolerance).

A variety of surface active properties of amino acids based surfactants like surface tension, CMC, phase behaviour and Krafft temperature have been studied by many scientists and it was concluded that amino acid surfactants have superior surface activity than their conventional counterparts. Properties of amino acid surfactants are discussed below:

7.1. Critical micelle concentration (CMC)

Critical micelle concentration (CMC) is an important phenomenon of surfactants as it governs a number of surface active properties such as solubilization, lytic action and their interaction with biological membranes etc. Commonly, increasing in the chain length of hydrocarbon tail (increasing hydrophobicity), resulted in decreased CMC values of surfactant solution and hence increased their surface activity (Mukerjee & Mysels, 1971). When compared to conventional surfactants, amino acid based surfactants generally have lower CMC values (Perez *et al.*, 2004).

Infante *et al.*, 1984 synthesized three arginine based amino acid surfactants by varying the combination of head group and hydrophobic tail (mono cationic amide, dicationic amide and di cationic amido-ester) and studied their CMC and gamma CMC values (surface tension at the CMC and the results showed a decrease in CMC and Gamma CMC values by increasing the length of hydrophobic tail. In a different study Mhatre & Singare, 2012 also

confirmed the CMC decreases with an increase in the number of carbon atoms in hydrophobic tail of N α -acyl arginine surfactants [Table1].

Yoshimura *et.al.*, 2012 investigated the CMC values of amino acid based gemini surfactants derived from cysteine and showed the decreasing pattern in CMC values by increasing the carbon length from 10 to 12 within the hydrophobic chain. Further increases in the carbon chain length upto 14 resulted in an increase in CMC values, confirming the low aggregation tendency of gemini surfactants.

Faustino *et al.*, 2011 reported the formation of mixed micelles in aqueous solutions of cystine based anionic gemini surfactants. The comparison has also been made with the gemini surfactant with its conventional monomeric counterpart, C₈Cys. CMC values for lipid-surfactant mixtures were reported to be lower than that of pure surfactant. The gemini surfactant and 1,2-diheptanoyl-sn-glycero-3-phosphocholine, a water-soluble, micelle-forming phospholipid was shown to have CMC values in the milimolar range.

Shrestha & Aramaki, 2009 studied the visco-elastic wormlike micelles formation in the aqueous systems of mixed type amino acid-based anionic and non-ionic surfactants in the absence of salt. In this investigation, N-dodecylglutamic acid was found to have higher Krafft temperature; however, when neutralized with alkaline amino acid L-lysine, it generated the micelles and the solution started behaving like a Newtonian fluid at 25°C.

7.2. Good water solubility

Good water solubility of amino acid surfactants is due to presence of additional CO-NH linkage (Moran *et al.*, 2004). This ability makes them readily biodegradable and environmentally friendly in comparison to their conventional counterparts. The presence of two carboxylic groups leads to even better solubility of N-acyl-L-glutamic acid in

1
2
3
4 water.Presence of two ionic arginine groups in one molecule of $C_n(CA)_2$ also exhibits good
5
6 water solubility thus resulting in more effective adsorption and diffusion on the cell
7
8 interface and showing effective antimicrobial action even at lower concentrations.
9

10 11 **7.3 Krafft temperature and Krafft point**

12
13
14 Krafft temperature can be understood as the unusual solubility behaviour of surfactants and
15
16 rapid increase in their solubilities above specific temperature (Pilemand, 2002). Ionic
17
18 surfactants have the affinity to produce solid hydrates through precipitation from the
19
20 aqueous medium. The solubility of the surfactants will commonly be observed to undergo a
21
22 sharp and discontinuous increase at some distinctive temperature, which is referred to as the
23
24 Krafft temperature. Krafft point of ionic surfactants is the Krafft temperature at their critical
25
26 micelle concentration (Malik & Ali, 2016).
27
28
29
30

31 This solubility characteristic is generally observed in ionic type surfactants and can
32
33 be explained on the basis of limited solubility of non-associated surfactants below Krafft
34
35 temperature and the solubility increasing gradually until the Krafft point is reached, which
36
37 was due to micelle formation. To achieve complete solubility it is necessary to prepare
38
39 surfactant formulations above their Krafft point (Puerto, 2001).
40
41
42

43 Krafft temperature of amino acids surfactants have been studied by various
44
45 researchers and comparison were made with the Krafft temperature of conventional
46
47 synthetic surfactants. Shrestha et. al., 2009 studied the Krafft temperature of arginine based
48
49 amino acid surfactants which were found to exhibit premicellar aggregation CMC values
50
51 above 2.5×10^{-6} M followed by the normal micellization at 3.6×10^{-4} . Ohta *et al.*, 2003
52
53 synthesized six different types of N-hexadecanoyl amino acid surfactants, and discussed
54
55 their relationship between the Krafft temperature and the amino acid residue.
56
57
58
59
60
61
62
63
64
65

1
2
3
4 In an experiment, an increase in the Krafft temperature of N-hexadecanoyl amino
5 acid surfactant was found with decrease in the size of the amino acid residue (phenylalanine
6 is an exceptional case). Whereas, an increase in the enthalpy of solution was found
7 endothermic with decrease in the size of the amino acid residue (except for glycine and
8 phenylalanine). From these observations it was concluded that D-L interaction was superior
9 to the L-L interaction in solid state of N-hexadecanoyl amino acid surfactant salt in both the
10 alanine system and the phenylalanine system.
11
12
13
14
15
16
17
18
19
20

21 Brito *et al.*, 2011 also determined the Krafft temperatures of three series of novel
22 amino acid based surfactants using differential scanning micro calorimetry and found the
23 change from trifluoroacetate ion to iodide which causes a relative large increase ($\approx 6^{\circ}\text{C}$) in
24 Krafft temperature from 47 to 53°C . Presence of a cis-double bond and unsaturation present
25 in long chain Ser-derivative significantly decreases the Krafft temperature. N-
26 Dodecylglutamic acid was reported to have higher Krafft temperature. However,
27 neutralization with alkaline amino acid L-lysine resulted in micelles formation in the
28 solution which behaves like a Newtonian fluid at 25°C .
29
30
31
32
33
34
35
36
37
38
39
40

41 **7.4. Surface tension**

42
43 Surface tension of surfactants increased with increasing chain length of the
44 hydrophobic moiety. Zhang *et al.*, 2013 determined the surface tension values of sodium
45 cocoylglycinate by Wilhelmy plate method using a DCAT11 tensiometer at $25^{\circ}\text{C} \pm 0.2^{\circ}\text{C}$
46 and obtained surface tension values of $33 \text{ mN} \cdot \text{m}^{-1}$ at the CMC of $0.21 \text{ mmol} \cdot \text{L}^{-1}$. Yoshimura
47
48
49
50
51
52
53
54
55
56
57
58
59
60
61
62
63
64
65 et.al., 2007 determined the surface tension of some 2CnCys type amino acid based
surfactants. The surface tension at CMC were found to lower when increasing the chain

length up to $n=8$, whereas the trend reversed for surfactants with $n=12$ or longer chain lengths.

The effect of CaCl_2 on surface tension of dicarboxylic amino acids type surfactants has also been studied (Bordes *et al.*, 2009). In this investigation addition of CaCl_2 in aqueous solutions of three dicarboxylic amino acid-based surfactants ($\text{C}_{12}\text{MalNa}_2$, $\text{C}_{12}\text{AspNa}_2$ and $\text{C}_{12}\text{GluNa}_2$) was carried out. Surface tensions were found to be lower at very low CaCl_2 compared to that at the plateau value beyond the CMC. This was due to the effect of calcium ions on the packing of the anionic surfactants at the air-water interface. Surface tension of the salts of N-dodecyl aminomalonic acid and N-dodecyl aspartic acid remained almost constant up to a 10 mM concentration of CaCl_2 . Above 10 mM a rapid increase in surface tension was observed, which was due to the formation of calcium salt precipitate of the surfactant. In case of disodium salt of N-dodecylglutamic acid, a considerable decrease in surface tension was observed by moderate addition of CaCl_2 that remained almost constant over a broad range of CaCl_2 concentrations.

For the determination of adsorption kinetics at air/water interface for gemini amino acid surfactants, dynamic surface tension measurements were carried out via the maximum bubble pressure method. Results revealed no change in the dynamic surface tension for $2\text{C}_{12}\text{Cys}$ for the longest measurement time. Reduction in the dynamic surface tension only depends on the length of hydrocarbon tail, number of hydrocarbon tails and the concentration. Faster decay was observed by increasing the concentration and decreasing the chain length and chain numbers of the surfactants. Results obtained at higher concentrations of C_nCys ($n=8-12$) were found very close to the γ_{cmc} obtained through Wilhelmy surface tension.

In other research (Fan *et al.*, 2008) the dynamic surface tension of SDLC (sodium dilauroylcystine) and SDDC (sodium didecaminocystine) was determined using Khan DCA-315 tensiometer through Wilhelmy plate technique and the equilibrium surface tension of their aqueous solution was measured by drop volume method. Further studies on the reaction of the disulfide bond were also performed by other methods. Addition of mercaptoethanol to 0.1 mmol L⁻¹ SDLC solution resulted in the rapid increase in surface tension from 34 to 53 mNm⁻¹. As NaClO can oxidize the disulfide groups of SDLC to sulfonic groups, when NaClO (5 mmol L⁻¹) was added into 0.1 mM L⁻¹SDLC solution. Transmission Electron Microscopy (TEM) and Differential Light Scattering (DLS) results showed no aggregate formation in the solution. The surface tension of SDLC was found to increase from 34 to 60 mNm⁻¹ over the duration of 20 min.

7.5 Binary surface interaction

The vibrant properties of the mixture of cationic type amino acid surfactants (diacylglycerol-arginine based surfactant) with phospholipids at the air-water interface for life sciences have been reported by many research groups and it was finally concluded this non-ideal characteristic as an asset to the pervasiveness of electrostatic interactions (Hines *et al.*, 1997; Wu *et al.*, 2017).

7.6. Aggregation properties

DLS measurements are commonly used to determine the aggregation properties of amino acid based monomeric and gemini surfactants, at concentrations higher than the CMC and the diffusion coefficient so obtained is converted into the apparent hydrodynamic diameter $D_H (=2R_H)$. When compared to the other surfactants, Cn Cys and $2Cn$ Cys formed relatively large aggregates with a wide size distribution. Surfactants other than $2C_{12}$ Cys

generally form aggregates of about 10 nm. The micelle size for gemini surfactants was found significantly larger than their monomeric counterparts (Faustino *et al.*, 2009). Increase in the hydrocarbon chain length also results in the increase micellular size (Brito *et al.*, 2009). Ohta *et al.*, 2007 described the aggregation properties of aqueous solutions of three different stereoisomers of tetramethyl ammonium N- dodecyl phenyl alanyl phenyl alaninate and revealed that the diastereomers have same critical aggregation concentrations in aqueous solutions. Iwahashi *et al.*, 2009 investigated the formation of chiral aggregates of optically active N-lauroyl-L-glutamic acid (L-LGA), N-lauroyl-L-valine (L-LVA) and its methyl ester in different solvents like tetrahydrofuran, acetonitrile, 1,4-dioxane and 1,2-dichloroethane through measurement of circular dichroism (CD), NMR chemical shift of N-H proton and vapour pressure osmometry.

7.7. Interfacial adsorption

Amino acids based surfactants have also been characterized in terms of their interfacial adsorption and compared to their conventional counterparts. Dodecyl ester of aromatic amino acids obtained from l-tyrosine (LET) and l-phenylalanine (LEP) were subjected to assess their interfacial adsorption characteristics (Vijay *et al.*, 2008). Results revealed that LET and LEP exhibit lower interfacial area at the air/solution and in water/hexane interface, respectively.

Bordes *et al.*, 2009 investigated the solution behaviour and adsorption of three dicarboxylic amino acid-based surfactants, the disodium salts of dodecylglutamic acid, dodecylaspartic acid and dodecylaminomalonic acid with three, two and one carbon atoms, respectively between the carboxyl groups at the air–water interface. It was reported that CMC of dicarboxylic surfactants was 4-5 times higher than mono carboxyl group containing

dodecylglycinate. This was attributed to the presence of hydrogen bonding between adjacent molecules through the amide groups in the dicarboxylic surfactants.

7.8. Phase behaviour

Phase behaviour of the surfactants can be determined using sealed ampoules containing requisite quantities of reagents which were mixed at higher temperatures and get homogenized with a vortex mixer through repeated centrifugation. After mixing properly, the samples are kept in a thermostatically controlled water bath at 25°C for a few days to ensure equilibrium. The equilibrated phases were acknowledged by visual observations using normal and crossed polarizer. At high concentration of surfactant, the isotropic discontinuous cubic phase was observed. A large hydrophilic head group of the surfactant molecule favoured the small and discrete aggregates of positive curvature (Johnsson & Edwards, 2001). Marques *et al.*, 2008 studied the phase behaviour of 12 Lys12/12 and 8 Lys8/16 Sersystems and stated that the 12 Lys12/12 Sersystem has phase separation between the micellar and vesicle solution. Small micelles were found to coexist in between the vesicles in the former whereas 8 Lys8/16 Sersystems showed a continuous transition process (where the elongated micelles were found to be present in between the small micelles and vesicles).

7.9. Emulsification

Kouchi *et al.*, 2001 examined the emulsification potency of N-[3-lauryl-2-hydroxypropyl]-L-arginine, L-glutamate, along with other amino acid surfactants, in terms of interfacial tension, dispersability and viscosity. Comparison of synthesized surfactants has also been made with their conventional non-ionic and amphoteric counterparts and

1
2
3
4 results revealed that the amino acid surfactants have greater emulsification potency than
5
6 conventional surfactants.
7

8
9 Baczko *et al.*, 2004 synthesized novel amino acid-based anionic surfactants and
10 investigated their usability as chiral oriented NMR spectroscopy solvents. A series of
11 sulfonated amphiphilic L-Phe or L-Ala derivatives with pentyl to tetradecyl tails were
12 synthesized by reacting the corresponding amino acid esters with o-sulfobenzoic anhydride.
13
14 Wu *et al.*, 2014 synthesized SFAAA (sodium N-fatty acyl amino acid) surfactant using pupa
15 oil and obtained PPH (pupa protein hydrolysates) as a waste by-product of the silk industry.
16
17 Studies on the emulsifying power of PPH and SFAAA in an oil/water emulsion revealed
18 that these surfactants were found superior with ethyl acetate as the oil phase over n-hexane.
19
20
21
22
23
24
25
26
27

28 **7.10. Lime tolerance**

29
30
31 Lime tolerance can be understood as the resistance of surfactant to precipitate as lime
32 soap in the presence of calcium, magnesium like ions present in hard water. Surfactants with
33 high tolerance against water hardness were found useful for detergent formulations and
34 personal care products. Determination of the lime tolerance can be achieved by calculating
35 the solubility and the surface activity change of surfactants in the presence of calcium ions
36 (Bordes *et al.*, 2009). Another method used to determine lime tolerance involve the
37 calculation of percentage or the number of grams of surfactant required for the dispersion of
38 lime soap formed from 100 g sodium oleate in water (Holmberg *et al.*, 2004). In hard water
39 areas, high concentrations of calcium and magnesium ions and mineral content can lead to
40 difficulties in some practical applications. Sodium ions are generally used as the counter
41 ions in the synthesis of anionic surfactants. By virtue of being bivalent calcium ions bind to
42 two surfactant molecules resulting in the easier surfactant precipitation from solution and
43
44
45
46
47
48
49
50
51
52
53
54
55
56
57
58
59
60
61
62
63
64
65

thus reducing detergency. Studies on lime tolerance of amino acid surfactants showed that 'Acid and lime tolerance' is vividly augmented by additional carboxyl group and they further increase with increase in the length of spacer group between two carboxyl groups increase. The order of acid and lime tolerance found was C₁₂Glycinate < C₁₂Aspartate < C₁₂Glutamate. Comparison of dicarboxylamido linkage surfactants with dicarboxyl amino linkage surfactants revealed that dicarboxyl amino linkage surfactants have wider pH range in which surface activity increases with moderate addition of acid. Dicarboxyl N-alkyl amino acids have a chelating effect in the presence of calcium ions and C₁₂Aspartate forms white gel. C₁₂Glutamate shows high surface activity at high concentration of Ca⁺⁺ and has the potential to get used in seawater desalination (Li, 2012).

7.11. Dispersability

Dispersability is the ability of surfactants to inhibit the agglomeration and settling down of surfactants in solution. Dispersability is an important property of surfactants that make them suitable for detergents, cosmetics and pharmaceuticals (Satyanarayana, 2012). Dispersing agents must contain an ester, ether, amido, or amino linkage between the hydrophobic group and the terminal hydrophilic group, as well as a straight-chain hydrophobic group (Rosen & Kunjappu, 2012). Generally the anionic surfactants like sulphated alkanolamide and zwitterionic surfactants such as amidosulfobetaine are found to be very effective as lime soap dispersing agents (Linfield, 1978; Weil *et al.*, 1970; Parris *et al.*, 1973; Parris *et al.*, 1976).

Various studies have been made to determine the dispersability of amino acid surfactants, in which N-lauroyllysine was found to have poor compatibility with water and was found difficult to make cosmetic formulations. In this series, N-acyl basic amino acid

was found to have superior dispersability and is being utilized in cosmetic industries to improve formulations (Sagawa *et al.*, 2015).

8. Toxicity

Conventional surfactants, especially cationic surfactants are intensely toxic against aquatic life which includes algae, fish and molluscs. Their acute toxicity is because of their tendency to disrupt the integral membrane which is due to the adsorption-ionic interaction phenomenon of surfactant at the cell water interface. Lowering the CMC of surfactants typically results in greater adsorption of surfactants onto interfaces, typically raising their acute toxicity. Increasing the length of hydrophobic chain of the surfactant molecule also results in the increase of acute toxicity of surfactant. Low or non-toxicity of the majority of amino acid surfactants to human and the environment, specifically to marine organisms, make them suitable as food ingredients, medicines and cosmetics (Ito & Inoue, 1982; Berger & Gacon, 1992; Davila *et al.*, 1997; George *et al.*, 1998; Myers, 2005; Obata *et al.*, 2008; Sagawa *et al.*, 2015). Various researchers have verified amino acid-based surfactants as mild and non-irritating to the skin (George *et al.*, 1998). Arginine based surfactants are known to be less toxic than their conventional counterparts. Brito *et al.*, 2009 studied the physicochemical and toxicological properties of amino acid-based amphiphiles and their spontaneously formed cationic vesicles derived using tyrosine (Tyr), hydroxyproline (Hyp), serine (Ser) and lysine (Lys and presented data for their acute toxicity to *Daphnia magna* (IC₅₀). In this research work they synthesized cationic vesicles of dodecyltrimethylammonium bromide (DTAB)/Lys-derivative and or Ser-/Lys-derivative mixtures and tested their ecotoxicity and hemolytic potential and stated that all amino acid surfactants along with their vesicle-containing mixtures showed lower ecotoxicity than the

1
2
3
4 conventionally used surfactant DTAB. Rosa *et al.*, 2007 studied the association of DNA and
5
6 stable cationic amino acid-based vesicles. As compared to conventionally used cationic
7
8 surfactants that are generally toxic, the interaction of cationic amino acid-based surfactants
9
10 appeared to be nontoxic. This cationic amino acid surfactant based on arginine, ALA, gives
11
12 spontaneously stable vesicles with some with anionic surfactants. Corrosion inhibitors based
13
14 on amino acids were also reported to be non-toxic by various researchers (Morad, 2008; de
15
16 Souza & Spinelli, 2009; Umoren *et al.*, 2009; Singh & Quraishi, 2010; Abiola & James,
17
18 2010; Barouni *et al.*, 2014). These surfactants were also found to be easy to synthesize with
19
20 high purities (up to 99%), low in cost, readily biodegradable and were completely soluble in
21
22 aqueous media. In various studies, sulfur containing amino acids based surfactants were
23
24 found to be more competent corrosion inhibitors (El-Naby, 1985; Rahim *et al.*, 1997;
25
26 Abiola, 2005; Morad, 2008; Özcan *et al.*, 2008; Chandra & Tyagi, 2013).
27
28
29
30
31
32

33 Perinelli *et al.*, 2017 in their latest investigation reported the favorable toxicological
34
35 profile of Rhamnolipids over conventional synthetic surfactants. Rhamnolipids were already
36
37 known as permeability enhancers. They have also reported the impact of Rhamnolipids on
38
39 the epithelial permeability of macromolecular drugs.
40
41
42

43 **9. Antimicrobial activity**

44
45 Antimicrobial activity of surfactants can be calculated on the basis of minimal
46
47 inhibitory concentration (MIC) (Infante *et al.*, 1984). Arginine based surfactants have been
48
49 studied in detail for their antimicrobial properties by many researchers (Franklin & Snow,
50
51 1981; Infante & Moses, 1994; Clapes *et al.*, 1999; Xia., 2001; Burczyk, 2003; Yoshimura *et*
52
53 *al.*, 2012). The gram negative bacteria were found to be more resistant to arginine based
54
55 surfactants than gram positive bacteria. Antimicrobial activity of surfactants typically
56
57
58
59
60
61
62
63
64
65

increases by the presence of hydroxyl group, cyclopropane or unsaturation within the acyl chain. Castillo *et al.*, 2006 stated that the length of acyl chains and the positive charge that determine the HLB of the molecule do affect their membrane disrupting property. $N\alpha$ -acylarginine methyl ester is another important class of cationic surfactants that are known to have broad spectrum antimicrobial activity with easy biodegradability and less or no toxicity (Kamimura, 1973; Piera *et al.*, 1998). Studies on the interaction of $N\alpha$ -acyl arginine methyl ester based surfactants with 1, 2-dipalmitoyl-*sn*-glycero-3-phosphocoline (DPPC) and 1, 2-dimiristoyl-*sn*-glycero-3-phosphocoline (DMPC), model membranes, as well as with living organisms with or without external barriers also revealed their good antimicrobial activity (Singare & Mhatre, 2012).

10. Biodegradability

Kamimura, 1973, Shida *et al.*, 1973 and 1975, and Kubo *et al.*, 1976 extensively studied the biodegradability of amino acid based surfactants and found that N-acylamino acids are easily biodegrade through decomposition into amino acid and fatty acids.

Single hydrophobic chain containing surfactants are comparatively more biodegradable than their branched counterpartssuch as Bis (Args). Typically, the more hydrophobic the surfactant the poorer will be their biodegradability. Akinari *et al.*, 2004synthesized the amino acid surfactants based on fatty acids rich amino acid surfactants and studied their physicochemical properties and biodegradability. Biodegradation studies of these surfactants showed their microbial degradation between 57-73% over 14 days. Zhang *et al.*, 2016 prepared supramolecular hydrogels mixtures of the biosurfactants based on sodium deoxycholate (NaDC) and amino acids such as glycine (Gly), alanine (Ala), lysine (Lys) and arginine (Arg) using different buffered solutions and claimed their unique

sensitivity towards multi-stimuli environments, their facile biodegradability and pH-sensitiveness makes them promising and versatile vehicles for dye (or drug) delivery.

11. Haemolytic activity

Nogueira *et al.*, 2011 investigated the five anionic lysine-based amino acid surfactants that differ on the basis of their counter ion and examined their ability to disrupt the cell membrane under varied pH range, concentrations and incubation period. For this purpose they used a standard hemolysis assay as a model for endosomal membranes. Results confirmed the pH sensitive hemolytic activity and better kinetics of these surfactants at the endosomal pH range.

Surfactants are known to have the capability to interact with lipid bilayer of cell membranes. R.B.C. (red blood cell) is one of the most frequently used cellular membranes as reference model to investigate the mechanisms of fundamental surfactant-induced osmotic cell resistance. Pérez *et al.*, 2009 studied the mechanisms of surfactant membrane interaction by monitoring the action of three arginine-based cationic amino acid surfactants and five lysine-based anionic amino acid surfactants on hypotonic hemolysis. Results revealed the dissimilar anti-haemolytic behaviours among amino acid-based surfactants, both linked to the maximal protective concentration. Physicochemical and structural properties of these compounds dictated the protection against hypotonic hemolysis. A good correlation was found in between the CMC and the concentrations of cationic surfactants that resulted in maximum protection against hypotonic hemolysis. In contrast, no correlation was observed for the anionic surfactants. Lysine based surfactants differ only in their counterions, this difference being responsible for their anti-hemolytic potency and the hemolytic activities.

1
2
3
4 Toxicological study revealed that the capability of arginine based monomeric and
5
6
7 gemini surfactants to disrupt the erythrocyte membranes depends is size as well as
8
9 hydrophobicity dependent (Tavano *et al.*, 2013).

10
11 Pinheiro & Faustino, 2017 discussed erythrocyte interaction with Na α , N ϵ -dioctyl
12
13 lysinate salts with different counterions (Li⁺, Na⁺, K⁺, Lys⁺, and Tris⁺). Surfactants
14
15 interacted with erythrocyte membranes in a biphasic way, at low concentrations through
16
17 protecting against hypotonic hemolysis and at high concentrations by inducing hemolysis.
18
19

20 21 **12. Rheological Properties**

22
23 Rheological properties of surfactants play a very important role to decide/predict
24
25 their applications in different industries such as food, pharmaceuticals, oil extraction and in
26
27 personal/home care products (Yang, 2002; Lin *et al.*, 2016; Kumar & Mandal, 2017). The
28
29 relation of viscoelastic characteristics and CMC of amino acid based surfactants have been
30
31 discussed by many researchers (Shrestha & Aramaki, 2009; Singh *et al.*, 2015; Saavedra *et*
32
33 *al.*, 2017).
34
35
36
37

38 39 **13. Industrial Applications**

40
41 Special structural characteristics, non-toxicity and biodegradability of AAS make
42
43 them suitable for various industrial applications.
44

45 46 **13.1. In agriculture**

47
48 Amino acid surfactant can be used as insecticides, herbicides and plant growth
49
50 inhibitors for agricultural means. Betaine ester surfactants are a class of cationic surfactants
51
52 that can be used as ‘temporary biocides’ and can be hydrolyzed very easily into harmless
53
54 components (Shida, *et al.*, 1973; Solans *et al.*, 1990; Tamarkin *et al.*, 2014).
55
56
57
58
59
60
61
62
63
64
65

1
2
3
4 A lawn pesticide was also reported in an US patent, that was synthesized using the
5
6 mixture of refined oils extracted from the plants of *Cupressaceae* family and an amino acid
7
8 derived surfactant solution, in which amino acid derived surfactant contributed 20 to 50%
9
10 by weight of the said solution. Herbicidal action of nonionic type amino acid-surfactant has
11
12 also been reported by various workers (Shida *et al.*, 1975).
13
14
15

16 **13.2. In laundry detergents**

17

18
19 Nowadays, the demand for amino acid based detergent formulations is increasing
20
21 globally. Amino acid surfactants are known to have better cleaning ability, foaming ability
22
23 and fabric softening properties that make them suitable to be used in house hold detergents,
24
25 shampoos, body wash etc. An amino acid surfactant derived from aspartic acid is reported as
26
27 an effective detergent with ampholytic and chelating characteristics. Less skin irritating
28
29 effects were observed by using detergent compositions made up of N-Alkyl- β -amino-
30
31 ethoxyacids (Leonard, 1976). Liquid detergent compositions made up of N-coco- β -
32
33 aminopropionate was claimed as an effective cleaning agent for grease mark on metallic
34
35 surfaces (Kennedy *et al.*, 1980; Cooper *et al.*, 1988). An aminocarboxylic surfactant,
36
37 $C_{14}CHOHCH_2NHCH_2COONa$, was also proved to have better detergency and used to clean
38
39 textiles, carpets, hair, glass etc (Miyamoto *et al.*, 1988). 2-hydroxy-3-aminopropionic acid
40
41 N,N di acetic acid derivatives were also known to incorporate stability in bleaching agents
42
43 with good complexing ability.
44
45
46
47
48
49

50
51 Detergent composition based on N-(N'-long chain acyl- β -alanyl)- β -alanine were
52
53 reported to have better ability and stability, foam breakage and good softening properties
54
55 (Miyamoto *et al.*, 1988). Keigo & Tatsuya, 1996 prepared and patented acyl amino acid
56
57 based detergent compositions. Kao formulated N-acyl-1-N-hydroxy- β -alanine based
58
59
60
61
62
63
64
65

detergent compositions and claimed their low skin irritation property, high water resistance and high detergency (Nozaki, 1995). Ajinomoto company used low toxic and easily degradable, L-glutamic acid, L-arginine and L-Lysine based amino acid surfactants as the principal ingredients in shampoos, detergents and cosmetics (Fig. 9). Ability of enzymatic additives in detergent compositions to remove protein based soilings has also been documented in the literature (Keshwani *et al.*, 2015). Glutamic acid, alanine, methyl glycine, serine and aspartic acid derived N-acyl amino acid surfactants were reported as good liquid cleaning agents in aqueous solutions. These surfactants did not show any increase in viscosity even at very low temperatures and can easily transfer from foamer container, thus providing uniform foam (Moriyama, 1998).

13.3. Lubricants

Amphoteric type amino acid surfactants are generally used as lubricants. Having excellent adhesive property for hydrophilic surfaces and low friction coefficient make them ideally suited to be lubricants. Various researchers prepared and studied the lubricating properties of AAS specifically glutamic acid (Jiang *et al.*, 2005; Jiang *et al.*, 2006; Xia & Jianhua, 2007).

13.4. In medicine

13.4.1. Drug carriers and preparation of functional liposomes

In the recent years, various researchers have claimed the ability of synthetic acyl amino acid/peptide as drug carrier and for preparing functional liposomes with lipopeptide ligands. As compared to the conventional liposomes of lecithin, vesicles of long aliphatic chain N α -acyl amino acids also showed encapsulation efficiencies for solutes (Boeckler *et al.*, 1998; Yagi *et al.*, 2000).

13.4.2. Gene therapy and DNA transfection

Gene therapy is a very important technique currently used in life science to safely introduce selected gene into living cells. Gemini surfactants have been found to have potential to be used as a carrier for the transportation of bioactive molecules. Lysine and 2, 4-diaminobutyric acid based polycationic gemini surfactants can be easily synthesized using standard peptide chemistry (Kirby *et al.*, 2003).

McGregor *et al.*, 2001 prepared a new class of amino acid based gemini surfactants as carrier for gene delivery into cells. Preliminary results revealed that combining these amino acid based gemini surfactants with dioleoylphosphatidylethanolamine (DOPE) allowed the synthesis of liposomes of different sizes and lipid compositions. Investigation revealed that suspensions of DOPE/surfactants mixtures in water, leads to a mixture of lipid vesicles with more complex structures corresponding to particles e500 nm in diameter. Different molar ratios of DOPE and surfactants (50/50, 60/40, and 70/30) were found to effect luciferase expression in CHO cells at comparable levels. The size of colloid and the molecular composition of the gemini surfactants matter in delivering optimum gene expression in living models.

DNA transfection efficiency of cysteine based amino acid surfactants with low molecular weight and its corresponding gemini were evaluated by Pena *et al.*, 2017. These surfactants showed no cytotoxicity and had greater efficacy than their commercially available counterparts to transfect CHO-K1 cells.

13.4.3. Antiviral agents

Lipoamino acids are attracting researchers due to their significant antiviral activity. Some acyl amino acid derivatives have also been reported to inhibit influenza

neuraminidase (Kondoh *et al.*, 2009). Several *N* α -palmitoylated amino acids/peptides when integrated into model membranes, affected the transition temperature between the surface tension at their critical micelle concentration (CMC).

Amino acid surfactants derived cationic surfactants obtained by condensation of fatty acids such as lauric acid and esterified dibasic amino acids such as Arginine may be exploited for the protection against the microorganism and these cationic surfactants were also found to be effective against viral infections. Furthermore, the addition of LAE into the cultures of Herpes virus type 1 Vaccinia virus and bovine parainfluenzae 3 virus leads to almost complete diminution of the virus organisms in these cultures, such effects being observed from 5 to 60 minutes (Bonvila *et al.*, 2006).

13.5. In food processing industry

Amino acid surfactants are widely used as emulsifiers in the food sectors such as in margarine, dairy products, low-calorie spreads and dressings (Flack, 1997). In addition, new insights into eating habits and health aspects need the changes in the food formulations time to time in terms of calories, fat content, vitamins or minerals that lead to a constant demand of optimizing product formulations. These surfactants also include optimal selection of raw materials. Formerly, Industries were giving great attention to the different glycerides, most commonly of monoglyceride derivatives but due to health consciousness, Arginine based amino acid surfactants with antimicrobial activity and wide antibacterial spectra are receiving great interest as a promising alternative of commercially available pure glycerides in food formulations. These surfactants additionally show considerable antimicrobial activity against *Escherichia coli* and *Salmonella*, which are the most common food borne pathogens and create a serious health hazard due to increased drug resistance (Lynde, 2001).

13.6. In cosmetic industry

AAS are used in the formulation of various personal care products. Potassium N-cocoylglycinate was found to be mild to skin and applied and used in face cleansers to remove soil and make-up. N-acyl-L-glutamic acid has two carboxylic groups, which lead to better water solubility. Among these, AAS based upon C₁₂ fatty acid are extensively used in face cleansers in order to clean soil and make up. AAS with C₁₈ have found to be used as emulsifiers in skin care cosmetics. N-dodecanoylalaninate is known to have ability to create non irritating creamy foam for skin, and therefore employed to formulate baby care products. N-lauroyl based AAS is used in toothpaste shows good detergency similar to soap and strong enzyme-inhibiting efficacy (Rosen & Kunjappu, 2012).

In the past few decades, in the selection of surfactants in cosmetics, personal care products and pharmaceuticals, less toxicity, mildness, gentleness and safety are becoming the major concerns. Consumers of these products are genuinely conscious for the potential irritation, toxicity and environmental factors. Nowadays, amino acids surfactants are used to formulate many shampoos, rinses, and body soaps because of the numerous merits over their conventional counterparts for use in cosmetics and personal care products. Protein based surfactants [PBS] possess desirable properties that a personal care product must have (Linfield, 1978). Some amino acid surfactants are known to have film forming ability whereas others have good foaming ability. Amino acids are important naturally occurring moisturizing factors (NMF) present in the stratum corneum. When epidermal cells die they become a part of the stratum corneum, and the proteins present in the cells gradually degrade into amino acids. These amino acids are then transported further into the stratum corneum, absorb the fat or fatlike substances into horny layer and thus improve the elasticity

1
2
3
4 of the skin surface. About 50% of the NMF present in the skin is made up of amino acids
5
6 and pyrrolidone (Cruz & Korchin,1994). Collagen, a common cosmetic ingredient also
7
8 comprises of amino acids, which keep the skin supple. Various skin problems such as
9
10 roughness and dark colouration of the skin are largely due to the deficiency of amino acids.
11
12 A case study revealed that mixing of an amino acid with ointment give relief to skin burn
13
14 and the affected area returned to the normal state and did not turn in to a keloid (Kiran *et al.*,
15
16 2010).
17
18
19
20

21 Amino acids are also found very useful for the care of damaged cuticles. Lacklustre
22
23 and dry hair can indicate a decrease of amino acid concentration in heavily damaged
24
25 cuticles. Amino acids have the ability to penetrate through the cuticle into hair shaft and
26
27 drawing in moisture from the skin. This capability of amino acid based surfactants make
28
29 them useful in shampoos, rinses, hair softeners, hair treatment agents, and the presence of
30
31 amino acids makes hair less breakable (Franklin & Snow, 1981).
32
33
34
35

36 **13.7. In microbial enhanced oil recovery**

37
38 Amino acid surfactants are also found useful in microbial enhanced oil recovery. A
39
40 bacterium strain, *B. aureum* MSA13 synthesized an amino acid surfactant that has potential
41
42 for use in microbial oil recovery. This surfactant has octadecanoic acid as a hydrophobic
43
44 moiety and hydrophilic group comprises of a tetrapeptide, a short sequence of four amino
45
46 acids pro-leu-gly-gly. Another example involve in microbial enhanced oil recovery in
47
48 marine surroundings is an actinobacterium, MSA13 produced surfactant (Zhou *et al.*, 2011).
49
50
51
52

53 **13.8. In nanomaterials**

54
55 Amino-acid-based polymerizable surfactants were also found to be useful in the
56
57 synthesis of chiral nanoparticles. Preiss et al., 2016 reported the synthesis of an amino-acid-
58
59
60
61
62
63
64
65

1
2
3
4 based chiral surfactants with polymerizable moieties which were then exploited to prepare
5
6 nanoparticles with a chiral surface functionality. They also checked their potential as
7
8 nucleating agents in the enantioselective crystallization of amino acid conglomerate systems,
9
10 taking a model system of rac-asparagine. Comparison were made in the particles synthesized
11
12 from chiral surfactants having different tail groups and results revealed that only the chiral
13
14 nanoparticles made up of polymerizable surfactant were found to have ability to act
15
16 efficiently as nucleation agent in the enantioselective crystallization.
17
18
19
20

21 **13.9. Other applications**

22
23 Amino acid surfactants are also known for use in forming PEDOT [Poly (3,4
24
25 ethylene dioxythiophene)] films (Hernandaz *et al.*, 2011; Moral-Vico *et al.*, 2013; Ahuja *et*
26
27 *al.*, 2007; Lowe , 1989; Malhotra *et al.*, 2006; Malhotra *et al.*, 2005; Andreescu & Sadik,
28
29 2004; Guimard *et al.*, 2007; Sackmann *et al.*, 1968; Sarfati *et al.*, 2000). Poly PEDOT films
30
31 through direct anodic oxidation of EDOT (3, 4-ethylenedioxythiophene) in aqueous solution
32
33 comprising of the sodium N-lauroylsarcosinate (SLS), an eco-friendly amino acid surfactant
34
35 were also prepared. In addition to the above mentioned applications, amino acid surfactants
36
37 are also used in the optimization of dry-cleaning process with carbon dioxide and as chiral
38
39 solvents (Sarfati *et al.*, 2001; Weiss-López *et al.*, 2001; Roosmalen *et al.*, 2004; Bordes &
40
41 Holmberg, 2015)
42
43
44
45
46
47

48 Roosmalen *et al.*, 2004 have optimized the dry-cleaning process using carbon
49
50 dioxide along with the amino acid surfactants, and studied the impact of various reaction
51
52 parameters on the cleaning. The reaction parameters studied were the cleaning time,
53
54 temperature, and power involve in mechanical action. In addition, the influence of the
55
56
57
58
59
60
61
62
63
64
65

1
2
3
4 quantity of each substance used such as CO₂, isopropyl alcohol, water and surfactant was
5
6 also investigated.
7

8
9 Self-assembly of amino acid surfactants resulted into micelles with a chiral surface.
10
11 Such a property of amino acid surfactant micelles of being supra-molecularly chiral makes
12
13 them suitable for asymmetric organic synthesis (Liu *et al.*, 2006). The chiral micelles can
14
15 also impart chirality in mesoporous materials prepared through surfactant templating route
16
17 and make their pores chiral. Dicarboxylic based amino acid surfactants can also behave like
18
19 surface-active chelating agents. It has been observed that the interaction of dicarboxylic
20
21 amino acid surfactants with divalent ion like calcium is reliant on the distance between the
22
23 two carboxyl groups. For example, N-acyl glutamate that has three –CH₂ groups between
24
25 the carboxyl groups does not form an intramolecular chelate with calcium, and is not able to
26
27 precipitate in water with high concentrations of calcium. The binding property of chelating
28
29 surfactants can also be used in mineral ore flotation. Calcium containing minerals like
30
31 calcite and apatite can be separated by means of flotation reagent with the right distance
32
33 between the carboxyl groups. N-alkyl amino acid surfactants are true amphoteric type
34
35 surfactants that can provide exceptionally low values of surface tension at the CMC, which
36
37 is due to formation of self-assemblies consisting of alternating anionic and cationic species.
38
39 This represents the micellization-driven protonation of surfactants.
40
41
42
43
44
45
46
47

48 Amino acid surfactants can also be used to design switchable surfactants. Cysteine
49
50 derivatives are the best known example of this class that can readily convert into cystine
51
52 derivatives via reversible processes. For example, long chain N-acylcysteine, a gemini
53
54 surfactant, is highly surface active and can be transformed into a cysteine derivative. Dithiothreitol, a
55
56 poor surface active agent, can also revert back to the gemini surfactant via an oxidation. The
57
58
59
60
61
62
63
64
65

1
2
3
4 exchange from one state to the other can also be achieved through electrochemical means
5
6 (Fiechter, 1992).
7
8

9 **14. Current Challenges for Amino acid Surfactants**

10
11 The economic feasibility of large scale production of amino acid surfactants is a
12 major issue to be surmounted. Biotechnological processes involved in their production are
13
14 not easily cost effective especially for the processes where bulk use of surfactants is
15
16 required such as petroleum and environmental application. Purification of substances is
17
18 another problem linked with these surfactants which is a needed for pharmaceutical,
19
20 cosmetic and food applications. Along with these disadvantages, denaturation and
21
22 dissociation of amino acid surfactants into their subunits and their activity are largely
23
24 affected by salt solutions.
25
26
27
28
29
30

31 Various researchers have suggested remedies to overcome these problems. The
32
33 overall cost cutting in their production can be achieved through the utilization of waste
34
35 substrate after combating their polluting effect. Development of efficient bioprocesses and
36
37 their successful optimization is also required, which include the optimization of the culture
38
39 conditions and cost-effective recovery processes for the maximum production and recovery
40
41 (Mohamed *et al.* 2017; Wang & Dado, 2017; Vecino *et al.*, 2017).
42
43
44
45

46 **15. Conclusions**

47
48 Although conceptually simple the process of combining amino acids with a
49
50 hydrophobic tail can by its very nature offer a variety of synthetic routes, this in turn leading
51
52 to many potential structures. Since the first studies on the synthesis of simple amino acid
53
54 surfactants in 1909 research has extended into the production of cationic, anionic, non-ionic
55
56 and amphoteric molecules, with detailed characterisation and assessment of physiochemical
57
58
59
60
61
62
63
64
65

properties performed throughout. Amino acid surfactants have been shown to have wide reaching applications over different industrial sectors, and the diversity in properties borne out of their potential for structural variety will in the future allow for this range of applications to grow further. Additionally, facile biodegradability and non-toxicity of amino acid surfactants can make them superior to their conventional synthetic counterparts when selected for the right application. Examples even exist whereby the chirality of amino acid surfactants has been exploited in use, such as imparting chirality into the surface of micelle template mesoporous materials. Amino acid surfactants have been reported to have a wide range of applications in different industrial sectors. Facile biodegradability and non-toxicity of amino acid surfactants can make them superior to their conventional synthetic counterparts when selected for the right application. The major challenges associated with amino acid surfactants are their high production cost and their challenging isolation to high purity. Various researchers are working to minimize these problems through selecting suitable renewable alternatives as substrate along with designing cost effective and scalable procedures. Amino acids surfactants have the potential to be widely accepted commercially in various industrial sectors in near future, especially so when the diversity of their structure and physiochemical properties are considered.

Acknowledgement

Divya Bajpai Tripathy acknowledges Department of Science and Technology, Government of India for financial support through grant no. SR/WOS-A/CS-1056/2014 (G).

References

Abdullah, N. H., Radiman, S., & Ghazali, Z. (2009). Study on Phase Behaviour of Sodium Lauroyl Sarcosinate, An Amino acid Based Surfactant *In Conference Proceedings Kimia Bersama UKM-ITB VIII 9-11 Jun 2009*, 370.

Abiola, O. K. (2005). Adsorption of methionine on mild steel. *Journal of the Chilean Chemical Society*, 50(4), 685-690.

Abiola, O. K., & James, A. O. (2010). The effects of Aloe vera extract on corrosion and kinetics of corrosion process of zinc in HCl solution. *Corrosion Science*, 52(2), 661-664.

Ahuja, T., Mir, I. A., & Kumar, D. (2007). Biomolecular immobilization on conducting polymers for biosensing applications. *Biomaterials*, 28(5), 791-805.

Akinari, A., Asakura, K., & Osanai, S. (2004). Synthesis and characterization of novel amphiphiles containing amino acid and carbohydrate. *Journal of surfactants and detergents*, 7(3), 297-303.

Al-Sabagh, A. M., Harding, D. R. K., Kandile, N. G., Badawi, A. M., & El-Tabey, A. E. (2009). Synthesis of some novel nonionic ethoxylated surfactants based on α -amino acids and investigation of their surface active properties. *Journal of Dispersion Science and Technology*, 30(3), 427-438.

Andreescu, S., & Sadik, O. A. (2004). Trends and challenges in biochemical sensors for clinical and environmental monitoring. *Pure and applied chemistry*, 76(4), 861-878.

Arima, K., Kakinuma, A., & Tamura, G. (1968). Surfactin, a crystalline peptidelipid surfactant produced by *Bacillus subtilis*: Isolation, characterization and its inhibition of fibrin clot formation. *Biochemical and biophysical research communications*, 31(3), 488-494.

1
2
3
4 Aubert, L., & Dussault, L. (2008). U.S. Patent Application No. 12/007,922.

5
6 B. Burczyk, (2003) Novel Surfactants; Preparation, Applications, and Biodegradability In Holmberg, K. (Ed.). (2003). *Novel surfactants: preparation applications and biodegradability, revised and expanded*, 114, Crc Press.

7
8
9
10
11
12
13
14 Baczko, K., Larpent, C., & Lesot, P. (2004). New amino acid-based anionic surfactants and their use as enantiodiscriminating lyotropic liquid crystalline NMR solvents. *Tetrahedron: Asymmetry*, 15(6), 971-982.

15
16
17
18
19
20
21 Barouni, K., Kassale, A., Albourine, A., Jbara, O., Hammouti, B., & Bazzi, L. (2014). Amino acids as corrosion inhibitors for copper in nitric acid medium: Experimental and theoretical study. *J. Mater. Environ. Sci*, 5(2), 456-463.

22
23
24
25
26
27
28 Berger C, Gacon P (1992) Preparation of N-acyl derivatives of amino acid mixtures obtained from cereal protein hydrolyzates for cosmetics. N° FR9221318

29
30
31
32
33 Boeckler, C., Frisch, B., & Schuber, F. (1998). Design and synthesis of thiol-reactive lipopeptides. *Bioorganic & medicinal chemistry letters*, 8(15), 2055-2058.

34
35
36
37
38 Bondi, S. (1909). Lipoprotein and the analysis of degenerative adiposis lipopeptides, their meaning, synthesis and characteristics (laurylglycin and laurylalanin). *Z Biochem*, 17, 543.

39
40
41
42
43
44
45 Bonvila, X. R., Roca, S. F., & Pons, R. S. (2006). *U.S. Patent Application No. 12/375,774*.

46
47
48
49
50
51 Bordes, R., Tropsch, J., & Holmberg, K. (2009). Counterion specificity of surfactants based on dicarboxylic amino acids. *Journal of colloid and interface science*, 338(2), 529-536.

1
2
3
4 Bordes, R., Tropsch, J., & Holmberg, K. (2009). Role of an amide bond for self-
5 assembly of surfactants. *Langmuir*, 26(5), 3077-3083.
6
7

8
9 Brito, R. O., Marques, E. F., & Silva, S. G. (2009). ML doVale, P. Gomes, MJ
10 Araújo, JE Rodriguez-Borges, MR Infante, MT Garcia, I. Ribosa, MP Vinardell, M.
11 Mitjans. *Colloids Surf. B*, 72, 80.
12
13
14

15
16 Brito, R. O., Marques, E. F., Silva, S. G., do Vale, M. L., Gomes, P., Araújo, M. J. &
17 Vinardell, M. P. (2009). Physicochemical and toxicological properties of novel amino acid-
18 based amphiphiles and their spontaneously formed catanionic vesicles. *Colloids and*
19 *Surfaces B: Biointerfaces*, 72(1), 80-87.
20
21
22
23
24

25
26 Brito, R. O., Silva, S. G., Fernandes, R. M., Marques, E. F., Enrique-Borges, J., & do
27 Vale, M. L. C. (2011). Enhanced interfacial properties of novel amino acid-derived
28 surfactants: Effects of headgroup chemistry and of alkyl chain length and
29 unsaturation. *Colloids and Surfaces B: Biointerfaces*, 86(1), 65-70.
30
31
32
33
34

35
36 Dana, V., Aurelia, P., Irina, E. C., & Mihai, C. C. (2011). Aspects regarding the
37 synthesis and surface properties of some glycine based surfactants. *Scientific Bulletin-*
38 *Universitatea Politehnica din Bucuresti, Series B*, 73(3), 147-154.
39
40
41
42

43
44 Diniz Rufino, R., Moura de Luna, J., de Campos Takaki, G. M., & Asfora Sarubbo,
45 L. (2014). Characterization and properties of the biosurfactant produced by *Candida*
46 *lipolytica* UCP 0988. *Electronic Journal of Biotechnology*, 17(1), 6-6.
47
48
49

50
51 Faustino, C. M., Calado, A. R., & Garcia-Rio, L. (2009). New urea-based surfactants
52 derived from α , ω -amino acids. *The Journal of Physical Chemistry B*, 113(4), 977-982.
53
54

55
56 Cantacuzene, D., & Guerreiro, C. (1989). Optimization of the papain catalyzed
57 esterification of amino acids by alcohols and diols. *Tetrahedron*, 45(3), 741-748.
58
59
60
61
62
63
64
65

1
2
3
4 Castillo, E. J., Gerson, S. H., & Han, W. W. (2000). *U.S. Patent No. 6,146,622*.
5
6 Washington, DC: U.S. Patent and Trademark Office.
7

8
9 Castillo, J. A., Infante, M. R., Manresa, À., Vinardell, M. P., Mitjans, M., & Clapés,
10
11 P. (2006). Chemoenzymatic synthesis and antimicrobial and haemolytic activities of
12
13 amphiphilic bis (phenylacetylarginine) derivatives. *ChemMedChem*, 1(10), 1091-1098.
14

15
16 Chandra, N., & Tyagi, V. K. (2013). Synthesis, properties, and applications of amino
17
18 acids based surfactants: a review. *Journal of Dispersion Science and Technology*, 34(6),
19
20 800-808.
21

22
23 Claffey, D. J., Meyer, J. D., Beauvais, R., Brandt, T., Shefter, E., Kroll, D. J. &
24
25 Manning, M. C. (2000). Long chain arginine esters: a new class of cationic detergents for
26
27 preparation of hydrophobic ion-paired complexes. *Biochemistry and Cell Biology*, 78(1),
28
29 59-65.
30
31

32
33 Clapés, P., & Rosa Infante, M. (2002). Amino acid-based surfactants: enzymatic
34
35 synthesis, properties and potential applications. *Biocatalysis and Biotransformation*, 20(4),
36
37 215-233.
38
39

40
41 Clapes, P., Morán, C., & Infante, M. R. (1999). Enzymatic synthesis of
42
43 arginine-based cationic surfactants. *Biotechnology and bioengineering*, 63(3), 333-343.
44

45
46 Cooper, L. A., Simon, J., & Wilson, D. A. (1988). *U.S. Patent No. 4,786,440*.
47
48 Washington, DC: U.S. Patent and Trademark Office.
49

50
51 Cruz, N. I., & Korchin, L. (1994). Inhibition of human keloid fibroblast growth by
52
53 isotretinoin and triamcinolone acetonide in vitro. *Annals of plastic surgery*, 33(4), 401-405.
54
55
56
57
58
59
60
61
62
63
64
65

1
2
3
4 Davila, A. M., Marchal, R., & Vandecasteele, J. P. (1997). Sophorose lipid
5 fermentation with differentiated substrate supply for growth and production phases. *Applied*
6 *microbiology and biotechnology*, 47(5), 496-501.

7
8
9
10
11 De Souza, F. S., & Spinelli, A. (2009). Caffeic acid as a green corrosion inhibitor for
12 mild steel. *Corrosion science*, 51(3), 642-649.

13
14
15
16 Castillo, E. J., Gerson S. H., & Han, W W. (2001) Anionic amino acid based
17 surfactants to enhance antimicrobial effectiveness of topical pharmaceutical compositions,
18
19 WO 2001021210 A1.

20
21
22
23 Soo, E. L., Salleh, A. B., Basri, M., Rahman, R. N. Z. R. A., & Kamaruddin, K.
24 (2003). Optimization of the enzyme-catalyzed synthesis of amino acid-based surfactants
25 from palm oil fractions. *Journal of bioscience and bioengineering*, 95(4), 361-367.

26
27
28
29 El-Naby, A., Khalil, N., & Mohamed, A. (1985). Inhibition by amino acids of the
30 corrosion of steel. *Surf Technol*, 24, 383-389.

31
32
33
34 Fan, H., Han, F., Liu, Z., Qin, L., Li, Z., Liang, D. & Fu, H. (2008). Active control of
35 surface properties and aggregation behavior in amino acid-based Gemini surfactant
36 systems. *Journal of colloid and interface science*, 321(1), 227-234

37
38
39
40 Faustino, C. M., Calado, A. R. & Garcia-Rio, L. (2011). Mixed micelle formation
41 between amino acid-based surfactants and phospholipids. *Journal of colloid and interface*
42 *science*, 359(2), 493-498.

43
44
45
46 Fiechter, A. (1992). Integrated systems for biosurfactant synthesis. *Pure and applied*
47 *chemistry*, 64(11), 1739-1743.

48
49
50
51 Flack, E. (1997). Butter, margarine, spreads, and baking fats. *Lipid Technologies and*
52 *Applications*. New York: Marcel Dekker, Inc, 305-327.

1
2
3
4 Fördedal, H., Sjöblom, J., & Infante, R. (1993). Lipoamino acid association in the
5 system N α -lauroyl-L-arginine methyl ester—1-pentanol—water as studied by dielectric
6 spectroscopy. *Colloids and Surfaces A: Physicochemical and Engineering Aspects*, 79(1),
7 81-88.
8
9

10
11
12
13
14 Franceschi, S., de Viguerie, N., Riviere, M., & Lattes, A. (1999). Synthesis and
15 aggregation of two-headed surfactants bearing amino acid moieties. *New Journal of*
16 *Chemistry*, 23(4), 447-452.
17
18
19

20
21 Franklin, T. J., & Snow, G. A. (1981). Biochemistry of Antibacterial Action.
22

23
24 Zhang, G., Xu, B., Han, F., Zhou Y., Liu, H., Li, Y., Cui, L., Tan, T. and Wang, N.
25 (2013) *Amer. J. Anal. Chem.*, 4, 2013.
26
27

28
29 Gallot, B., & Hassan, H. H. (1989). Lyotropic lipo-amino-acids: synthesis and
30 structural study. *Molecular Crystals and Liquid Crystals*, 170(1), 195-214.
31
32

33
34 George, A., Modi, J., Jain, N., & Bahadur, P. (1998). A Comparative Study on the
35 Surface Activity and Micellar Behaviour of Somen-Acylamino acid Based
36 Surfactants. *Indian journal of chemistry. Sect. A: Inorganic, physical, theoretical &*
37 *analytical*, 37(11), 985-992.
38
39
40
41

42
43 Gerecht, J. F. (1977). *U.S. Patent No. 4,048,338*. Washington, DC: U.S. Patent and
44 Trademark Office.
45
46

47
48 Gerova, M., Rodrigues, F., Lamère, J. F., Dobrev, A., & Fery-Forgues, S. (2008).
49 Self-assembly properties of some chiral N-palmitoyl amino acid surfactants in aqueous
50 solution. *Journal of colloid and interface science*, 319(2), 526-533.
51
52
53

54
55 Godtfredsen, S. E., & Bjoerkling, F. (1990). An enzyme-catalyzed process for
56 preparing N-acyl amino acids and N-acyl amino acid amides. *World Patent*, (90/14429).
57
58
59
60
61
62
63
64
65

1
2
3
4 Guimard, N. K., Gomez, N., & Schmidt, C. E. (2007). Conducting polymers in
5 biomedical engineering. *Progress in polymer science*, 32(8), 876-921.
6
7

8
9 Hathout, Y., Ho, Y. P., Ryzhov, V., Demirev, P., & Fenselau, C. (2000). Kurstakins:
10 A New Class of Lipopeptides Isolated from *Bacillus thuringiensis*. *Journal of natural*
11 *products*, 63(11), 1492-1496.
12
13

14
15 Hentrich, W., Keppler, H., & Hintzmann, K. German Pat. 635,522 (Sept. 18, 1936);
16 IG Farbenind, A.-G. *British Pats*, 459(461,328).
17
18

19
20 Hines, J. D., Thomas, R. K., Garrett, P. R., Rennie, G. K., & Penfold, J. (1997).
21 Investigation of mixing in binary surfactant solutions by surface tension and neutron
22 reflection: anionic/nonionic and zwitterionic/nonionic mixtures. *The Journal of Physical*
23 *Chemistry B*, 101(45), 9215-9223.
24
25

26
27 Holmberg, K., Jönsson, B., Kronberg, B., & Lindman, B. (2004). Surfactants and
28 polymers in aqueous solution. *Journal of Synthetic Lubrication*, 20(4), 367-370.
29
30

31
32 <https://www.ulprospector.com/en/eu/PersonalCare/Suppliers/6835/Ajinomoto-Omni>
33 Chem.
34
35

36
37 Husmann, M. (2008). Surfactants Derived from Naturally-Occurring Amino
38 acids. *SÖFW-Journal*, 134(3).
39
40

41
42 Infante, M. R., Mollnero, J., & Erra, P. (1992). Lipopeptidic surfactants. II. Acidic
43 and basic N α -lauroyl-L-arginine dipeptides from pure amino acids. *Journal of the American*
44 *Oil Chemists Society*, 69(7), 647-652.
45
46

47
48 Infante, M., & Moses, V. (1994). Synthesis and surface activity properties of
49 hydrophobic/hydrophilic peptides. *International journal of peptide and protein*
50 *research*, 43(2), 173-179.
51
52
53
54
55
56
57
58
59
60
61
62
63
64
65

1
2
3
4 Infante, M., Pinazo, A., & Seguer, J. (1997). Non-conventional surfactants from
5 amino acids and glycolipids: structure, preparation and properties. *Colloids and Surfaces A:*
6
7 *Physicochemical and Engineering Aspects*, 123, 49-70.
8
9

10
11 Infante, R., Dominguez, J. G., Erra, P., Julia, R., & Prats, M. (1984). Surface active
12 molecules: preparation and properties of long chain α -acyl-1- α -amino- ω -guanidine alkyl
13 acid derivatives. *International Journal of Cosmetic Science*, 6(6), 275-282.
14
15
16

17
18 Ito, S., & Inoue, S. (1982). Sophorolipids from *Torulopsis bombicola*: possible
19 relation to alkane uptake. *Applied and Environmental Microbiology*, 43(6), 1278-1283.
20
21
22

23
24 Ivanković, T., & Hrenović, J. (2010). Surfactants in the environment. *Arhiv za*
25 *higijenu rada i toksikologiju*, 61(1), 95-109.
26
27

28
29 Iwahashi, M., Matsuzawa, H., Minami, H., Yano, T., Wakabayashi, T., Ino, M., &
30 Sakamoto, K. (2002). Solvent effect on chiral aggregate formation of acylamino
31 acids. *Journal of Oleo Science*, 51(11), 705-713.
32
33
34

35
36 Johnsson, M., & Edwards, K. (2001). Phase behavior and aggregate structure in
37 mixtures of dioleoylphosphatidylethanolamine and poly (ethylene glycol)-
38 lipids. *Biophysical Journal*, 80(1), 313-323.
39
40
41

42
43 Kamimura, A. (1973). Colorimetric Determination of Long-chain N-Acylglutamic
44 Acids with Pinacyanol. *Agricultural and Biological Chemistry*, 37(3), 457-464.
45
46

47
48 Kango, N. (2010). *Textbook of microbiology*. IK International Pvt Ltd.
49

50
51 Keigo S. and Tatsuya H., Detergent composition (1996) *U. S. 5529712 A*.
52

53
54 Kennedy, R. R., Lindemann, M. K., & Verdicchio, R. J. (1980). *U.S. Patent No.*
55 *4,181,634*. Washington, DC: U.S. Patent and Trademark Office.
56
57
58
59
60
61
62
63
64
65

1
2
3
4 Keshwani, A., Malhotra, B., & Kharkwal, H. (2015). Natural polymer based
5
6 detergents for stain removal. *World J Pharm Pharm Sci*, 4(4), 490-508.
7

8
9 Kim, P. I., Ryu, J., Kim, Y. H., & Chi, Y. T. (2010). Production of biosurfactant
10
11 lipopeptides Iturin A, fengycin and surfactin A from *Bacillus subtilis* CMB32 for control of
12
13 *Colletotrichum gloeosporioides*. *J Microbiol Biotechnol*, 20(1), 138-145.
14
15

16 Kimura, Y., Kobayashi, Y., Adachi, S., & Matsuno, R. (1992). Aminoacylase-
17
18 Catalyzed synthesis of N-acyl amino acid from fatty acid or its ethyl ester and amino acid.
19
20 In *Biochemical Engineering for 2001* (pp. 109-111). Springer Japan.
21
22

23
24 Kiran, G. S., Thomas, T. A., Selvin, J., Sabarathnam, B., & Lipton, A. P. (2010).
25
26 Optimization and characterization of a new lipopeptide biosurfactant produced by marine
27
28 *Brevibacterium aureum* MSA13 in solid state culture. *Bioresource technology*, 101(7),
29
30 2389-2396.
31
32

33 Kirby, A. J., Camilleri, P., Engberts, J. B., Feiters, M. C., Nolte, R. J., Söderman, O.
34
35 & Guédat, P. (2003). Gemini surfactants: new synthetic vectors for gene
36
37 transfection. *Angewandte Chemie International Edition*, 42(13), 1448-1457.
38
39

40 Kondoh, M., Furutani, T., Azuma, M., Ooshima, H., & Kato, J. (1997). Acyl amino
41
42 acid derivatives as novel inhibitors of influenza neuraminidase. *Bioscience, Biotechnology,*
43
44 *and Biochemistry*, 61(5), 870-874.
45
46

47 Kouchi, J., Tabohashi, T., Yokoyama, S., Harusawa, F., Yamaguchi, A., Sakai, H.,
48
49 & Abe, M. (2001). Emulsifying Potency of New Amino acid-Type Surfactant (1). O/W
50
51 Emulsions. *Journal of oleo science*, 50(11), 847-855.
52
53
54
55
56
57
58
59
60
61
62
63
64
65

1
2
3
4 Kubo, M., Yamada, K., & Takinami, K. (1976). Effects of chemical structure on
5
6 biodegradation of long chain N-acyl amino acids. *Journal of Fermentation Technology*
7
8 (*Japan*).

9
10
11 Kumar, S., & Mandal, A. (2017). Rheological properties and performance evaluation
12
13 of synthesized anionic polymeric surfactant for its application in enhanced oil
14
15 recovery. *Polymer*, 120, 30-42.

16
17
18 Kunieda, H., Nakamura, K., Infante, M. R., & Solans, C. (1992). Reversed vesicles
19
20 from biocompatible surfactants. *Advanced Materials*, 4(4), 291-293.

21
22
23 Hernandez, L., Pingarron J.M. and Yanez-sedeno, P. In da la Guardia, M., &
24
25 Garrigues, S. (2011). *Challenges in green analytical chemistry* (No. 13). Royal Society of
26
27 Chemistry, 214.

28
29
30 Pérez, L., Pinazo, A., García, M. T., Lozano, M., Manresa, A., Angelet, M., &
31
32 Infante, M. R. (2009). Cationic surfactants from lysine: synthesis, micellization and
33
34 biological evaluation. *European journal of medicinal chemistry*, 44(5), 1884-1892.

35
36
37 Sánchez, L., Martínez, V., Infante, M. R., Mitjans, M., & Vinardell, M. P. (2007).
38
39 Hemolysis and antihemolysis induced by amino acid-based surfactants. *Toxicology*
40
41 *Letters*, 169(2), 177-184.

42
43
44 Leonard, E. O. (1976). *U.S. Patent No. 3,960,742*. Washington, DC: U.S. Patent and
45
46 Trademark Office.

47
48
49 Li, Y.X. (2011). Synthesis and physicochemical study of novel amino acid based
50
51 surfactants. Division of Applied Surface Chemistry Chalmers University of Technology,
52
53 Göteborg, Sweden.

1
2
3
4 Linfield, W. M. (1978). Soap and lime soap dispersants. *Journal of the American Oil*
5
6
7 *Chemists' Society*, 55(1), 87-92.

8
9 Liu, Y., Jessop, P. G., Cunningham, M., Eckert, C. A., & Liotta, C. L. (2006).
10
11 Switchable surfactants. *Science*, 313(5789), 958-960.

12
13
14 Lowe, C. R. (1989). Biosensors. *Philosophical Transactions of the Royal Society of*
15
16 *London B: Biological Sciences*, 324(1224), 487-496.

17
18
19 Lundberg, D., Faneca, H., Morán, M. D. C., Pedroso De Lima, M. C., Miguel, M. D.
20
21 G., & Lindman, B. (2011). Inclusion of a single-tail amino acid-based amphiphile in a
22
23 lipoplex formulation: effects on transfection efficiency and physicochemical
24
25 properties. *Molecular Membrane Biology*, 28(1), 42-53.

26
27
28 Malhotra, B. D., Chaubey, A., & Singh, S. P. (2006). Prospects of conducting
29
30 polymers in biosensors. *Analytica Chimica Acta*, 578(1), 59-74.

31
32
33 Malhotra, B. D., Singhal, R., Chaubey, A., Sharma, S. K., & Kumar, A. (2005).
34
35 Recent trends in biosensors. *Current Applied Physics*, 5(2), 92-97.

36
37
38 Marques, E. F., Brito, R. O., Silva, S. G., Rodríguez-Borges, J. E., Vale, M. L. D.,
39
40 Gomes, P., & Söderman, O. (2008). Spontaneous vesicle formation in cationic mixtures
41
42 of amino acid-based surfactants: chain length symmetry effects. *Langmuir*, 24(19), 11009-
43
44 11017.

45
46
47
48 McGregor, C., Perrin, C., Monck, M., Camilleri, P., & Kirby, A. J. (2001). Rational
49
50 approaches to the design of cationic gemini surfactants for gene delivery. *Journal of the*
51
52 *American Chemical Society*, 123(26), 6215-6220.

1
2
3
4 Mitin, Y. V., Braun, K., & Kuhl, P. (1997). Papain catalyzed synthesis of glyceryl
5
6 esters of N-protected amino acids and peptides for the use in trypsin catalyzed peptide
7
8 synthesis. *Biotechnology and Bioengineering*, 54(3), 287-290.
9

10
11 Miyamoto, N., Ikeuchi, T., & Shinjo, Z. (1988). *U.S. Patent No. 4,749,515*.
12
13 Washington, DC: U.S. Patent and Trademark Office.
14

15
16 Mohr, B., Boeckh, D., Norenberg, R., Randall, S., Panandiker, R., Gosselink, E. P.,
17
18 & Luipold, L. (2000). *U.S. Patent No. 6,034,204*. Washington, DC: U.S. Patent and
19
20 Trademark Office.
21

22
23 Lynde, C. W. (2001). Moisturizers: what they are and how they work. *Skin Therapy*
24
25 *Lett*, 6(13), 3-5.
26

27
28 Montet, D., Servat, F., Pina, M., Graille, J., Galzy, P., Arnaud, A., & Marcou, L.
29
30 (1990). Enzymatic synthesis of N- ϵ -acyllysines. *Journal of the American Oil Chemists'*
31
32 *Society*, 67(11), 771-774.
33

34
35 Morad, M. S. (2005). Effect of amino acids containing sulfur on the corrosion of
36
37 mild steel in phosphoric acid solutions containing Cl⁻, F⁻ and Fe³⁺ ions: behavior under
38
39 polarization conditions. *Journal of Applied Electrochemistry*, 35(9), 889-895.
40

41
42 Morad, M. S. (2008). Inhibition of iron corrosion in acid solutions by Cefatrexyl:
43
44 Behaviour near and at the corrosion potential. *Corrosion Science*, 50(2), 436-448.
45

46
47 Moral-Vico, J., Carretero, N. M., Pérez, E., Suñol, C., Lichtenstein, M., & Casañ-
48
49 Pastor, N. (2013). Dynamic electrodeposition of amino acid-polypyrrole on amino acid-
50
51 PEDOT substrates: Conducting polymer bilayers as electrodes in neural
52
53 systems. *Electrochimica Acta*, 111, 250-260.
54
55
56
57
58
59
60
61
62
63
64
65

1
2
3
4 Moran, C., Infante, M. R., & Clapes, P. (2001). Synthesis of glycerol amino acid-
5 based surfactants. Part 1. Enzymatic preparation of rac-1-O-(N α -acetyl-L-aminoacyl) glycerol
6 derivatives. *Journal of the Chemical Society, Perkin Transactions 1*, (17), 2063-2070.
7
8

9
10
11 Moran, M. C., Pinazo, A., Perez, L., Clapes, P., Angelet, M., García, M. T. &
12 Infante, M. R. (2004). "Green" amino acid-based surfactants. *Green Chemistry*, 6(5), 233-
13 240.
14
15
16

17
18
19 Moriyama, M., Tanabe, H., Hanazawa, H., & Kajihara, Y. (1998). *U.S. Patent No.*
20 *5,712,232*. Washington, DC: U.S. Patent and Trademark Office.
21
22

23
24 Mukerjee, P., & Mysels, K. J. (1971). *Critical micelle concentrations of aqueous*
25 *surfactant systems* (No. NSRDS-NBS-36). National Standard reference data system.
26
27

28
29 Myers, D. (2005). *Surfactant science and technology*. John Wiley & Sons.
30

31
32 Nogueira, D. R., Mitjans, M., Infante, M. R., & Vinardell, M. P. (2011). The role of
33 counterions in the membrane-disruptive properties of pH-sensitive lysine-based
34 surfactants. *Acta biomaterialia*, 7(7), 2846-2856.
35
36
37

38
39 Nagao, A., & Kito, M. (1989). Synthesis of O-acyl-L-homoserine by lipase. *Journal*
40 *of the American Oil Chemists' Society*, 66(5), 710-713.
41
42

43
44 Nozaki, T. (1995). *U.S. Patent No. 5,417,875*. Washington, DC: U.S. Patent and
45 Trademark Office.
46
47

48
49 Obata, Y., Suzuki, D., & Takeoka, S. (2008). Evaluation of cationic assemblies
50 constructed with amino acid based lipids for plasmid DNA delivery. *Bioconjugate*
51 *chemistry*, 19(5), 1055-1063.
52
53
54

55
56 Ohta, A., Ozawa, N., Nakashima, S., Asakawa, T., & Miyagishi, S. (2003). Krafft
57 temperature and enthalpy of solution of N-acyl amino acid surfactants and their racemic
58
59
60
61
62
63
64
65

1
2
3
4 modifications: effect of the amino acid residue. *Colloid & Polymer Science*, 281(4), 363-
5
6 369.
7

8
9 Ohta, A., Shirai, M., Asakawa, T., & Miyagishi, S. (2007). Effect of stereochemistry
10 on the molecular aggregation of phenylalanine dipeptide-type surfactants. *Journal of oleo*
11 *science*, 57(12), 659-667.
12
13
14

15
16 Ongena, M., Jourdan, E., Adam, A., Paquot, M., Brans, A., Joris, B., ...& Thonart, P.
17 (2007). Surfactin and fengycin lipopeptides of *Bacillus subtilis* as elicitors of induced
18 systemic resistance in plants. *Environmental microbiology*, 9(4), 1084-1090.
19
20
21

22
23 Özcan, M., Karadağ, F., & Dehri, I. (2008). Investigation of adsorption
24 characteristics of methionine at mild steel/sulfuric acid interface: An experimental and
25 theoretical study. *Colloids and Surfaces A: Physicochemical and Engineering*
26 *Aspects*, 316(1), 55-61.
27
28
29
30
31

32
33 Heitmann, P. (1968). Reactivity of Sulfhydryl Groups in Micelles. *European Journal*
34 *of Biochemistry*, 5(3), 305-315.
35
36

37
38 Heitmann, P. (1968). A model for sulfhydryl groups in proteins. Hydrophobic
39 interactions of the cysteine side chain in micelles. *The FEBS Journal*, 3(3), 346-350.
40
41

42
43 Jiang, H., Tao, D., & Wang, B. (2005, January). Synthesis of the Multifunctional
44 Additive N-acyl Glutamic Acid and Application in Water-Based Lubricating Fluids.
45 In *World Tribology Congress III* (pp. 667-668). *American Society of Mechanical Engineers*.
46
47
48

49
50 Jiang, H., Tao, D., Wang, B., Huang, S., Xu, J., & Qian, L. (2006). Synthesis and
51 Behavior of Friction and Wear of N-oleoylglutamic Acid as Water-based Lubricant
52 Additive. *Tribology Beijing*, 26(1), 48.
53
54
55
56
57
58
59
60
61
62
63
64
65

1
2
3
4 Lin, Z., Zheng, Y., Talmon, Y., Maxson, A., & Zakin, J. L. (2016). Comparison of
5 the effects of methyl-and chloro-substituted salicylate counterions on drag reduction and
6 rheological behavior and micellar formation of a cationic surfactant. *Rheologica Acta*, 55(2),
7 117-123.
8
9

10
11
12
13
14 Pang, X., & Chu, C. C. (2010). Synthesis, characterization and biodegradation of
15 functionalized amino acid-based poly (ester amide) s. *Biomaterials*, 31(14), 3745-3754.
16
17

18
19 Parris, N., Weil, J. K., & Linfield, W. M. (1973). Soap based detergent formulations.
20
21 V. Amphoteric lime soap dispersing agents. *Journal of the American Oil Chemists*
22 *Society*, 50(12), 509-512.
23
24

25
26 Parris, N., Weil, J. K., & Linfield, W. M. (1976). Soap-based detergent formulations:
27 XVIII. Effect of structure variations on surface-active properties of sulfur containing
28 amphoteric surfactants. *Journal of the American Oil Chemists Society*, 53(3), 97-100.
29
30

31
32 Pathak, K. V., & Keharia, H. (2014). Identification of surfactins and iturins produced
33 by potent fungal antagonist, *Bacillus subtilis* K1 isolated from aerial roots of banyan (*Ficus*
34 *benghalensis*) tree using mass spectrometry. *3 Biotech*, 4(3), 283-295.
35
36

37
38 Pegiadou, S., Perez, L., & Infante, M. R. (2000). Synthesis, characterization and
39 surface properties of 1-N-l-tryptophan-glycerol-ether surfactants. *Journal of surfactants and*
40 *detergents*, 3(4), 517-525.
41
42

43
44 Pérez, L., Infante, M. R., Pons, R., Morán, C., Vinardell, P., Mitjans, M., & Pinazo,
45 A. (2004). A synthetic alternative to natural lecithins with antimicrobial properties. *Colloids*
46 *and Surfaces B: Biointerfaces*, 35(3), 235-242.
47
48
49
50
51
52
53
54
55
56
57
58
59
60
61
62
63
64
65

Pérez, L., Pinazo, A., García, M. T., del Carmen Moran, M., & Infante, M. R. (2004). Monoglyceride surfactants from arginine: synthesis and biological properties. *New journal of chemistry*, 28(11), 1326-1334.

Pérez, L., Pinazo, A., Pons, R., & Infante, M. (2014). Gemini surfactants from natural amino acids. *Advances in colloid and interface science*, 205, 134-155.

Peypoux, F., Bonmatin, J. M., & Wallach, J. (1999). Recent trends in the biochemistry of surfactin. *Applied Microbiology and Biotechnology*, 51(5), 553-563.

Piera, E., Comelles, F., Erra, P., & Infante, M. R. (1998). New alkyl amide type cationic surfactants from arginine. *Journal of the Chemical Society, Perkin Transactions 2*, (2), 335-342.

Piera, E., Infante, M. R., & Clapés, P. (2000). Chemo-enzymatic synthesis of arginine-based gemini surfactants. *Biotechnology and bioengineering*, 70(3), 323-331.

Pilemand, C. (2002). Surfactants: Their abilities and important physico-chemical properties. *Arbejdsmiljøinstituttet, København*.

Peña, L. C., Argaraña, M. F., De Zan, M. M., Giorrello, A., Antuña, S., Prieto, C. C., ...& Müller, D. M. (2017). New Amphiphilic Amino acid Derivatives for Efficient DNA Transfection in Vitro. *Advances in Chemical Engineering and Science*, 7(02), 191.

Perinelli, D. R., Vllasaliu, D., Bonacucina, G., Come, B., Pucciarelli, S., Ricciutelli, M., ...& Casettari, L. (2017). Rhamnolipids as epithelial permeability enhancers for macromolecular therapeutics. *European Journal of Pharmaceutics and Biopharmaceutics*, 119, 419-425.

Pinazo, A., Manresa, M. A., Marques, A. M., Bustelo, M., Espuny, M. J., & Perez, L. (2016). Amino acid-based surfactants: new antimicrobial agents. *Advances in colloid and interface science*, 228, 17-39.

Pinheiro, L., & Faustino, C. (2017). Amino acid-Based Surfactants for Biomedical Applications. In *Application and Characterization of Surfactants*. InTech.

Polidori, A., Wathier, M., Fabiano, A. S., Olivier, B., & Pucci, B. (2006). Synthesis and aggregation behaviour of symmetric glycosylated bolaamphiphiles in water. *Arkivoc*, 4, 73-89.

Pinazo, A., Pons, R., Pérez, L., & Infante, M. R. (2011). Amino acids as raw material for biocompatible surfactants. *Industrial & Engineering Chemistry Research*, 50(9), 4805-4817.

Presenz, P. (1996). Lipoamino acids and lipopeptides as amphiphilic compounds. *Pharmazie*, 51(10), 755-758.

Preiss, L. C., Wagner, M., Mastai, Y., Landfester, K., & Muñoz-Espí, R. (2016). Amino acid based Polymerizable Surfactants for the Synthesis of Chiral Nanoparticles. *Macromolecular Rapid Communications*, 37(17), 1421-1426.

Puerto, M. C. (2001). *Surfactants: Fundamentals and Applications in the Petroleum Industry*: Cambridge University Press, 2000, pp. 621, £ 85.00 (US \$140.00)(hardback), ISBN 0-521-64067-9.

R. Bordes and K. Holmberg, *Adv. Colloid Interfac.*, 2015, 222, 79-91.

R. Bordes and K. Holmberg, Amino acid-based surfactants, In *Encyclopedia of Surface and Colloid Science*, 2nd edition; P. Somasundaran, Ed., Taylor and Francis: UK, 2004.

1
2
3
4 Rahim, M. A., Hassan, H. B., & Khalil, M. W. (1997). Naturally Occuring Organic
5
6 Substances as Corrosion Inhibitors for mild steel in acid medium. *Materialwissenschaft und*
7
8 *Werkstofftechnik*, 28(2), 98-102.
9

10
11 Rentsch, K. M. (2002). The importance of stereoselective determination of drugs in
12
13 the clinical laboratory. *Journal of biochemical and biophysical methods*, 54(1), 1-9.
14
15

16 Rosa, M., del Carmen Morán, M., da Graça Miguel, M., & Lindman, B. (2007). The
17
18 association of DNA and stable catanionic amino acid-based vesicles. *Colloids and Surfaces*
19
20 *A: Physicochemical and Engineering Aspects*, 301(1), 361-375.
21
22

23 Rosen, M. J., & Kunjappu, J. T. (2012). *Surfactants and interfacial phenomena*. John
24
25 Wiley & Sons.
26
27

28 Saavedra, L. C. C., Gómez, E. M. P., Oliveira, R. G., & Fernández, M. A. (2017).
29
30 Aggregation behaviour and solubilization capability of mixed micellar systems formed by a
31
32 gemini lipoamino acid and a non-ionic surfactant. *Colloids and Surfaces A:*
33
34 *Physicochemical and Engineering Aspects.*, 2(3), 1144–1148.
35
36
37

38 Sackmann, E., Meiboom, S., & Snyder, L. C. (1968). Nuclear magnetic resonance
39
40 spectra of enantiomers in optically active liquid crystals. *Journal of the American Chemical*
41
42 *Society*, 90(8), 2183-2184.
43
44

45 Sagawa, K., Gesslein, B. W., Popova, K., Oshimura, E., Ikeda, N., & Kamidoi, T.
46
47 (2015). *U.S. Patent No. 9,034,924*. Washington, DC: U.S. Patent and Trademark Office.
48
49

50 Sarfati, M., Aroulanda, C., Courtieu, J., & Lesot, P. (2001). Enantiomeric
51
52 recognition of chiral invertomers through NMR in chiral oriented phases: a study of cis-
53
54 decalin. *Tetrahedron: Asymmetry*, 12(5), 737-744.
55
56
57
58
59
60
61
62
63
64
65

1
2
3
4 Sarfati, M., Lesot, P., Merlet, D., & Courtieu, J. (2000). Theoretical and experimental
5 aspects of enantiomeric differentiation using natural abundance multinuclear NMR
6 spectroscopy in chiral polypeptide liquid crystals. *Chemical Communications*, (21), 2069-
7 2081.
8
9

10
11
12
13
14 Satyanarayana, T., Johri, B. N., & Prakash, A. (Eds.). (2012). *Microorganisms in*
15 *sustainable agriculture and biotechnology*. Springer Science & Business Media.
16
17

18
19 Seguer^a, J., Infante, M. R., Allouch, M., Mansuy, L., Selve, C., & Vinardell, P.
20 (1994). Synthesis and evaluation of non-ionic amphiphilic compounds from amino acids:
21 molecular mimics of lecithins. *New journal of chemistry*, 18(6), 765-774.
22
23

24
25
26 Seguer^b, J., Molinero, J., Manresa, A., & Caelles, J. (1994). Physicochemical and
27 antimicrobial properties of N"-acyl-L-arginine dipeptides from. *J. Soc. Cosmet. Chem*, 45,
28 53-63.
29
30

31
32
33 Shida, T., Homma, Y., & Misato, T. (1973). Bacterial Degradation of N-Lauroyl-L-
34 valine. *Agricultural and Biological Chemistry*, 37(5), 1027-1033.
35
36

37
38 Shida, T., Homma, Y., Kamimura, A., & Misato, T. (1975). Studies on the control of
39 plant diseases by amino acid derivatives. V. Degradation of N-lauroyl-L-valine in soil and
40 the effect of sunlight and ultraviolet rays on N-lauroyl-L-valine. *Agricultural and*
41 *Biological Chemistry*, 39(4), 879-883.
42
43

44
45
46 Shrestha, R. G., & Aramaki, K. (2009). The Study of Salt Induced Viscoelastic
47 Wormlike Micelles in Aqueous Systems of Mixed Anionic/Nonionic Surfactants. *Journal of*
48 *Nepal Chemical Society*, 23, 65-73.
49
50

51
52
53 Singare, P. U., & Mhatre, J. D. (2012). Cationic surfactants from arginine: synthesis
54 and physicochemical properties. *American Journal of Chemistry*, 2(4), 186-190.
55
56
57
58
59
60
61
62
63
64
65

1
2
3
4 Singh, A. K., & Quraishi, M. A. (2010). Effect of Cefazolin on the corrosion of mild
5 steel in HCl solution. *Corrosion Science*, 52(1), 152-160.
6
7

8
9 Singh, J., Michel, D., Getson, H. M., Chitanda, J. M., Verrall, R. E., & Badea, I.
10 (2015). Development of amino acid substituted gemini surfactant-based mucoadhesive gene
11 delivery systems for potential use as noninvasive vaginal genetic
12 vaccination. *Nanomedicine*, 10(3), 405-417.
13
14
15
16
17

18
19 Solans, C., Infante, R., Azemar, N., & Wörnheim, T. (1989). Phase behavior of
20 cationic lipoamino acid surfactant systems. In *Trends in Colloid and Interface Science*
21 *III* (pp. 70-75). Steinkopff.
22
23
24

25
26 Solans, C., Pés, M. A., Azemar, N., & Infante, M. R. (1990). Lipoamino acid
27 surfactants: Phase behavior of long chain N α -Acyl arginine methyl esters. In *Trends in*
28 *Colloid and Interface Science IV* (pp. 144-150). Steinkopff.
29
30
31
32

33
34 Takehara, M. (1989). Properties and applications of amino acid based
35 surfactants. *Colloids and Surfaces*, 38(1), 149-167.
36
37

38
39 Tamarkin, D., Eini, M., Friedman, D., Berman, T., & Schuz, D. (2014). *U.S. Patent*
40 *No. 8,795,693*. Washington, DC: U.S. Patent and Trademark Office.
41
42

43
44 Taubeneck, U. (1977). TJ FRANKLIN and GA SNOW, Biochemistry of
45 Antimicrobial Action. XII, 224 S., 85 Abb., 1 Tab. London 1975: Chapman and Hall.
46
47 \$39.50. *Zeitschrift für Allgemeine Mikrobiologie*, 17(4), 327-327.
48
49

50
51 Tavano, L., Infante, M. R., Riya, M. A., Pinazo, A., Vinardell, M. P., Mitjans, M., &
52 Perez, L. (2013). Role of aggregate size in the hemolytic and antimicrobial activity of
53 colloidal solutions based on single and gemini surfactants from arginine. *Soft Matter*, 9(1),
54
55
56
57
58 306-319.
59
60
61
62
63
64
65

Trivedi, T. J., Rao, K. S., Singh, T., Mandal, S. K., Sutradhar, N., Panda, A. B., & Kumar, A. (2011). Task-Specific, Biodegradable Amino acid Ionic Liquid Surfactants. *ChemSusChem*, 4(5), 604-608.

Ueda, N. (2003). *U.S. Patent No. 6,583,087*. Washington, DC: U.S. Patent and Trademark Office.

Umoren, S. A., Obot, I. B., & Obi-Egbedi, N. O. (2009).Raphia hookeri gum as a potential eco-friendly inhibitor for mild steel in sulfuric acid. *Journal of Materials Science*, 44(1), 274-279.

Valivety, R., Gill, I. S., & Vulfson, E. N. (1998). Application of enzymes to the synthesis of amino acid-based bola and gemini surfactants. *Journal of Surfactants and Detergents*, 1(2), 177-185.

Valivety, R., Jauregi, P., Gill, I., & Vulfson, E. (1997).Chemo-enzymatic synthesis of amino acid-based surfactants. *Journal of the American Oil Chemists' Society*, 74(7), 879-886.

Van Roosmalen, M. J. E., Woerlee, G. F., & Witkamp, G. J. (2004). Amino acid based surfactants for dry-cleaning with high-pressure carbon dioxide. *The Journal of supercritical fluids*, 32(1), 243-254.

Vecino, X., Cruz, J. M., Moldes A. B. & Rodrigues, L. R. (2017). Biosurfactants in cosmetic formulations: trends and challenges. *Critical Reviews in Biotechnology*, 37(7), 911-923.

Vijay, R., Angayarkanny, S., & Baskar, G. (2008). Amphiphilic dodecyl ester derivatives from aromatic amino acids: Significance of chemical architecture in interfacial

adsorption characteristics. *Colloids and Surfaces A: Physicochemical and Engineering Aspects*, 317(1), 643-649.

Vonderhagen, A., Rath, H. C., & Eilers, E. (1999). Enzymatic catalyzed N-acylation of amino acids, protein hydrolysates and/or their derivatives. *Ger. Offen.DE*, 19.

Walia, N. K., & Cameotra, S. S. (2015). Lipopeptides: Biosynthesis and applications. *J Microb Biochem Technol*, 7(2), 103-107.

Walther, W., & Netscher, T. (1996). Design and development of chiral reagents for the chromatographic ee determination of chiral alcohols. *Chirality*, 8(5), 397-401.

Wang, S. F., Furuno, T., & Cheng, Z. (2001). Synthesis of new amino acid-type amphoteric surfactants from tall oil fatty acid. *Journal of Wood Science*, 47(6), 470-475.

Wang, B. & Dado, G.P. (2017). Process for preparing N-acyl amino acid salts. US Patent 9,593,072, 2017

Weil, J. K., Parris, N., & Stirton, A. J. (1970). Synthesis and properties of sulfated alkanolamides. *Journal of the American Oil Chemists' Society*, 47(3), 91-93.

Weiss-López, B. E., Azocar, M., Montecinos, R., Cassels, B. K., & Araya-Maturana, R. (2001). Differential Incorporation of l- and d-N-Acyl-1-phenyl-d 5-2-aminopropane in a Cesium N-Dodecanoyl-l-alaninate Cholesteric Nematic Lyomesophase. *Langmuir*, 17(22), 6910-6914.

Winkelmann, G., Allgaier, H., Lupp, R., & Jung, G. (1983). Iturin AL--a new long chain iturin possessing an unusual high content of C16-beta-amino acids. *The Journal of antibiotics*, 36(11), 1451-1457.

Wu, M. H., Wan, L. Z., & Zhang, Y. Q. (2014). A novel sodium N-fatty acyl amino acid surfactant using silkworm pupae as stock material. *Scientific reports*, 4, 4428.

1
2
3
4 Wu, R., Qiu, X., Shi, Y., & Deng, M. (2017). Molecular dynamics simulation of the
5
6 atomistic monolayer structures of N-acyl amino acid-based surfactants. *Molecular*
7
8 *Simulation*, 43(7), 491-501.
9

10
11 Xia, G. J. C. B. S., & Jianhua, F. (2007). Study on Tribological Properties of N-
12
13 Acylglutamic Acid in Mineral Base Oil. *Lubrication Engineering*, 12, 023.
14
15

16 Xia, J. (2001). Protein-Based Surfactants: Synthesis: Physicochemical Properties,
17
18 and Applications (Vol. 101). CRC Press.
19
20

21 Xia, J., Xia, Y., & Nnanna, I. A. (1995). Structure-function relationship of acyl
22
23 amino acid surfactants: surface activity and antimicrobial properties. *Journal of agricultural*
24
25 *and food chemistry*, 43(4), 867-871.
26
27

28 Yagi, N., Ogawa, Y., Kodaka, M., Okada, T., Tomohiro, T., Konakahara, T., &
29
30 Okuno, H. (2000). Preparation of functional liposomes with peptide ligands and their binding
31
32 to cell membranes. *Lipids*, 35(6), 673-680.
33
34

35 Yang, J. (2002). Viscoelastic wormlike micelles and their applications. *Current*
36
37 *opinion in colloid & interface science*, 7(5), 276-281.
38
39

40 Yoshimura, T., Sakato, A., & Esumi, K. (2012). Solution properties and
41
42 emulsification properties of amino acid-based gemini surfactants derived from
43
44 cysteine. *Journal of oleo science*, 62(8), 579-586.
45
46

47 Yoshimura, T., Sakato, A., Tsuchiya, K., Ohkubo, T., Sakai, H., Abe, M., & Esumi,
48
49 K. (2007). Adsorption and aggregation properties of amino acid-based N-alkyl cysteine
50
51 monomeric and N, N'-dialkyl cystine gemini surfactants. *Journal of colloid and interface*
52
53 *science*, 308(2), 466-473.
54
55
56
57
58
59
60
61
62
63
64
65

1
2
3
4 Zhang, M., Strandman, S., Waldron, K. C., & Zhu, X. X. (2016). Supramolecular
5
6 hydrogelation with bile acid derivatives: structures, properties and applications. *Journal of*
7
8 *Materials Chemistry B*, 4(47), 7506-7520.
9

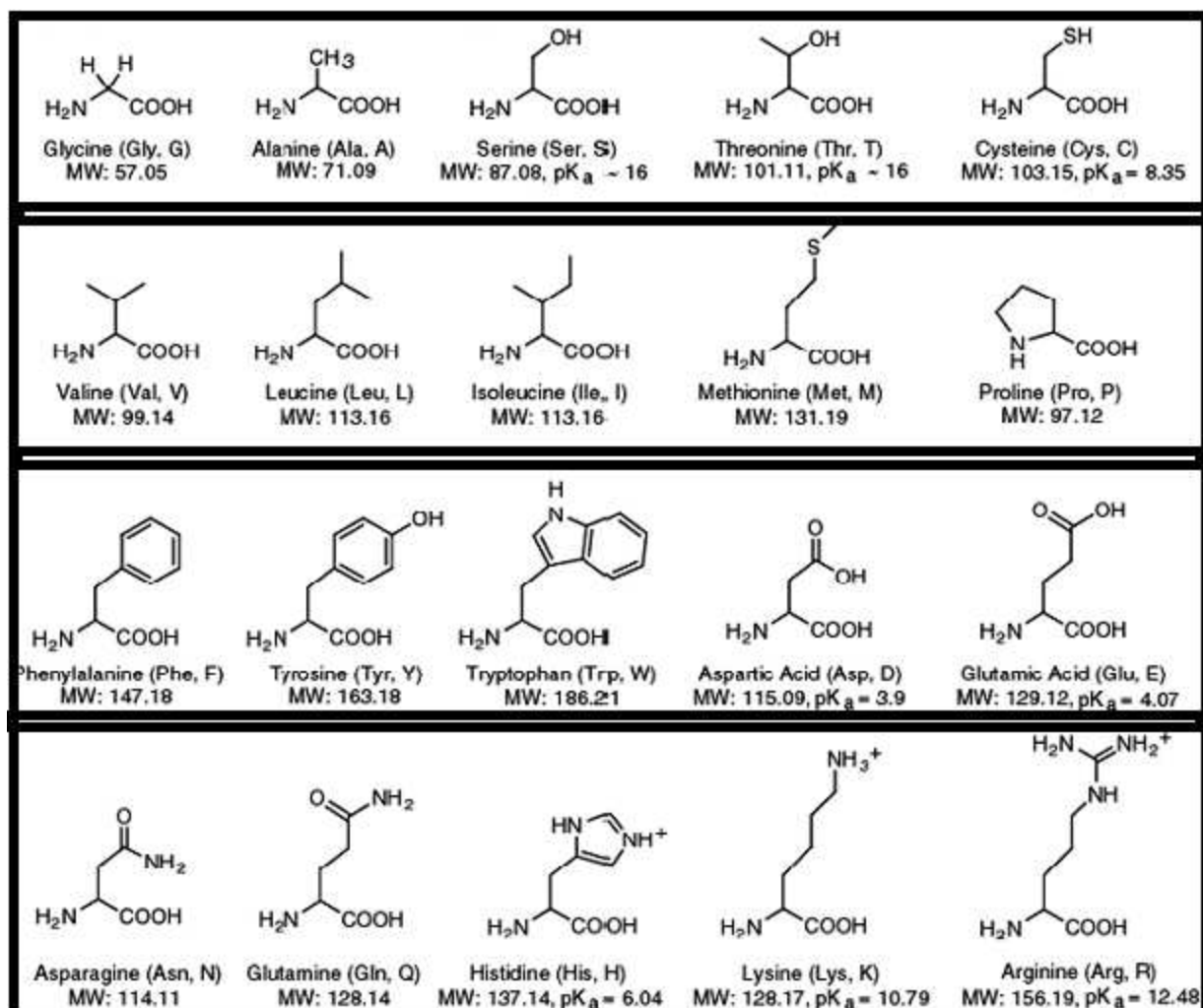
10
11 Zhou, C., Liu, Z., Yan, Y., Du, X., Mai, Y. W., & Ringer, S. (2011). Electro-
12
13 synthesis of novel nanostructured PEDOT films and their application as catalyst
14
15 support. *Nanoscale research letters*, 6(1), 364.
16
17

18
19 [https://www.ulprospector.com/en/eu/PersonalCare/Suppliers/6835/Ajinomoto-](https://www.ulprospector.com/en/eu/PersonalCare/Suppliers/6835/Ajinomoto-OmniChem.)
20
21 OmniChem.)
22
23
24
25
26
27
28
29
30
31
32
33
34
35
36
37
38
39
40
41
42
43
44
45
46
47
48
49
50
51
52
53
54
55
56
57
58
59
60
61
62
63
64
65

	Reviewers' Comments	Response
1	AAS applications (regarding several technological fields) are dependent on their toxicological profile. Authors should address in a more extensive and careful way, the toxicity issue and its relationship with the physicochemical properties of AAS. Recent papers like Perinelli et al, 2016 (Eur J Pharm Biopharm) could be added to the references list.	Section 8, Toxicity has been rewritten and incorporated all the suggestions and references given by the learned reviewer.
2	The same applies to the antimicrobial activity. There is currently a very broad database on AAS biomedical applications. It is mandatory to incorporate more up-to-date information. Examples of more recent papers are Tavano et al, 2013 (Soft Matter), Pinazo et al, 2016 (ACIS), and Coronel-Léon et al, 2017 (CSBB).	Section 9, Antimicrobial activity has been rewritten and all the suggestions and references given by the learned reviewer have been incorporated.
3	Did the authors considered the inclusion of a rheological appreciation of AAS? Besides Shrestha, R. G., &Aramaki, K. (2009), recent papers dealing with this subject seem quite appealing, since thickening, viscosity consistency and foaming are technological challenges related to many AAS industrial applications.	Rheological properties of AAS have been incorporated as a separate section (Section 12).
4	In terms of application in Medicine concerning drug delivery, the authors did not address the fact that AAS can be promising novel biomaterials in nano-particulate systems. Why?	It has been included in the manuscript in Section 12.8.
5	- Page 33, 12.4.2.: DNA transfection deserves also a more full approach. Literature reports, others than Claffe et al, 2000, are available. Furthermore, gene therapy could be included in DNA transfection item. Considering DNA transfection and Gene therapy, authors should rewrite the corresponding text in a more careful way, accurately relating the scientific results (e.g. what was the main achievement, in terms of transfection efficiency, after combining the Gemini AAS with DOPE?)	Gene therapy and DNA transfection have clubbed and the section 12.4.2. is dedicated to it.

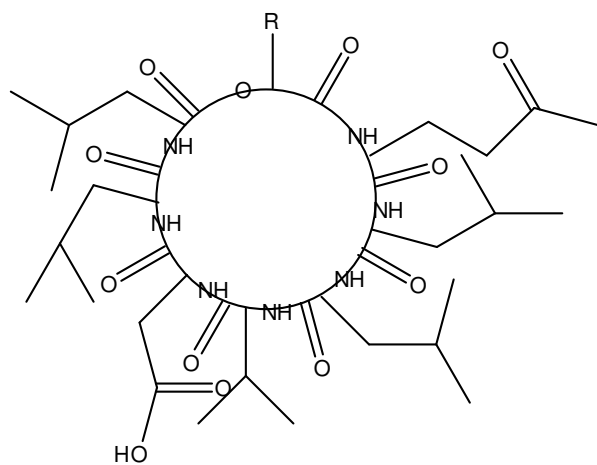
<p>size for gemini surfactants was significantly larger than their monomeric counterparts." Any citation?</p> <p>- Page 23, line -3: rephrase "In an another study (Bordes et al.,2009) investigated the solution behavior and adsorption of three dicarboxylic amino acid-based surfactants....."</p> <p>- Page 24, 7.8: More details (although briefly) and references are needed for this part.</p> <p>- Page 29, 10: Change the order to: "Kamimura, 1973, Shida et al., 1973 & 1975 and Kubo et al., 1976 extensively....."</p> <p>- Page 30, line -12.: the reference "145" is not in the same style of referencing as the others.</p> <p>- Page 30, 11: a more comprehensible explanation is needed for the interaction surfactant-erythrocyte membrane.</p> <p>- Page 33, 12.4.2.: DNA transfection deserves also a more full approach. Literature reports, others than Claffe et al, 2000, are available.</p> <p>- Figure 1: a more careful use of capital letters must be taken into account (an example is the caption of Figure 1).</p> <p>- Figure 1: the information about the meaning of this pKa is missing. pKa presented here is the value of side chain.</p> <p>- Figure 3: should be presented in a more careful manner.</p> <p>Figures 4, 5, 6 and 7: better structures representations are needed.</p> <p>- Figure 7: authors should use always the same notation throughout the text (e.g. "bolaamphiphiles" or "bola amphiphiles").</p> <p>- Figure 10 should be better presented, in order to address the corresponding caption.</p>	<p>Correction has been made</p> <p>Correction has been made</p> <p>Correction has been made</p> <p>Correction has been made</p> <p>Explanation has been added as per reviewer suggestions</p> <p>Corrections have been made accordingly</p> <p>Corrections have been made accordingly</p> <p>Figure 3 has been deleted</p> <p>Corrections have been made</p> <p>Corrections have been made</p> <p>Figure 10, which is now Figure 9 has been redrawn.</p>
---	--

Figure

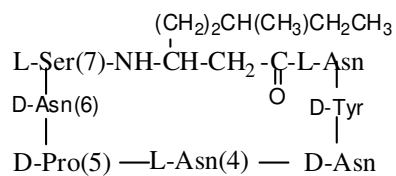
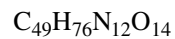


pK_a is the negative base-10 logarithm of the acid dissociation constant (K_a) of a solution. pK_a = -log₁₀K_a

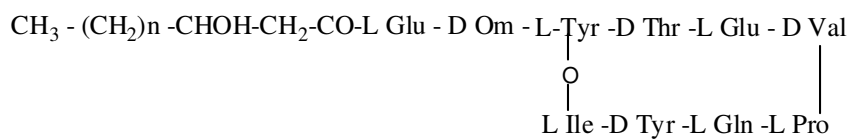
Figure 1: 20 standard aminoacids



a



b



c

Figure 2: a=surfactin, b=iturin A3 and c=fengycin B

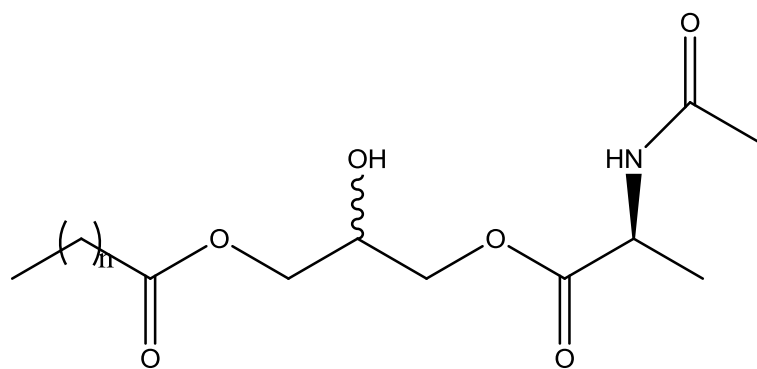


Figure 3: Example of linear AAS

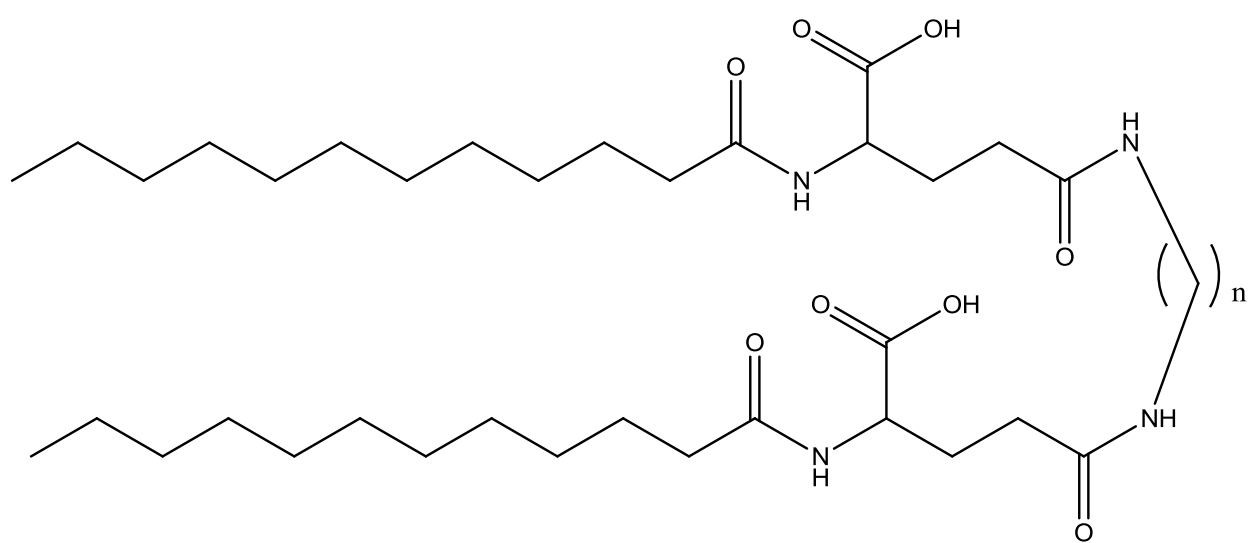


Figure 4: Example of gemini AAS ($s = 2, 5$, and 8).

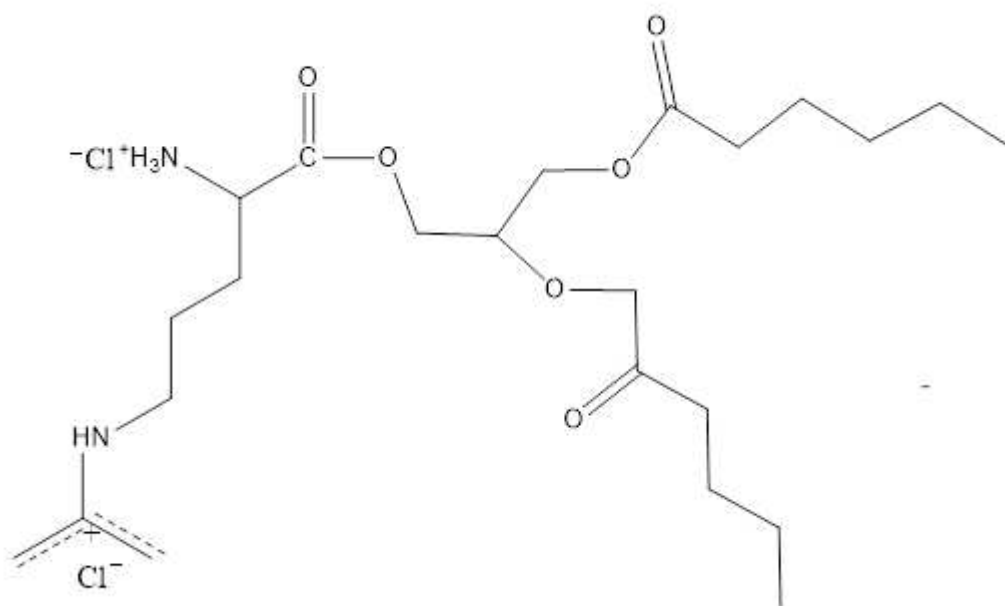


Figure 5: Example of glycerolipid type AAS

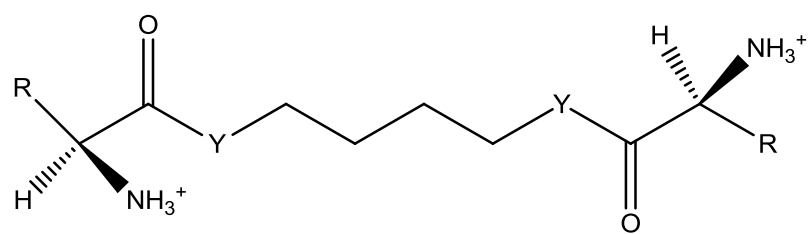


Figure6: Example of bolaamphiphiles AAS

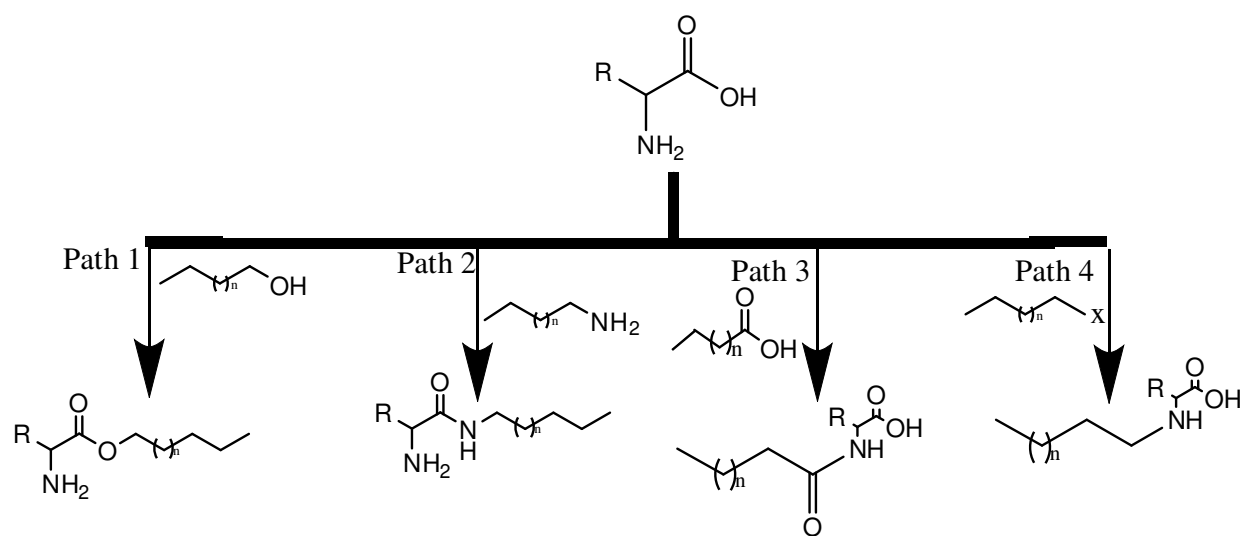


Figure 7: Fundamental synthesis paths of aminoacid based surfactants

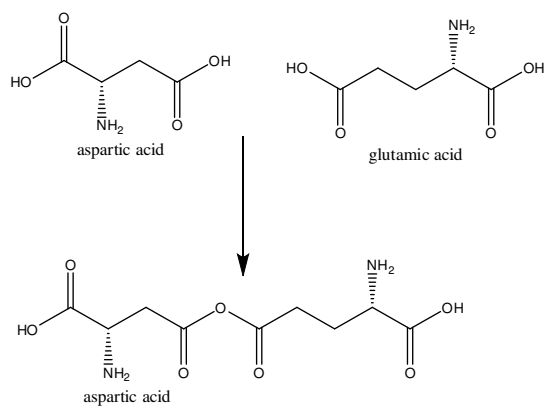


Figure8: Path 5 for synthesis of AAS

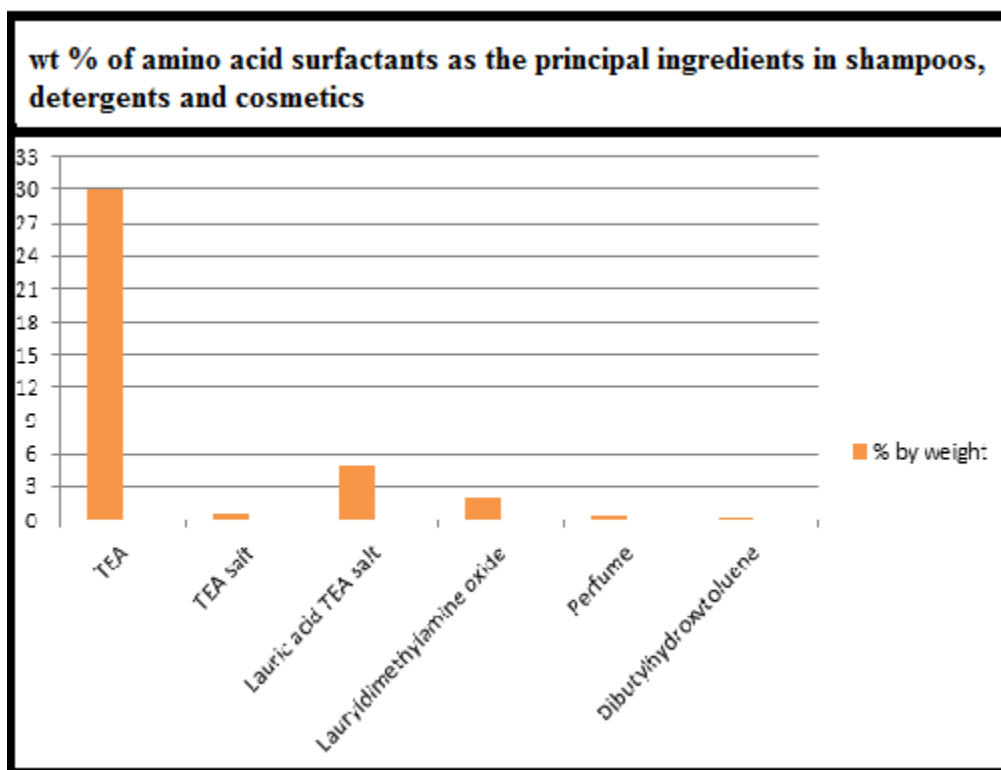
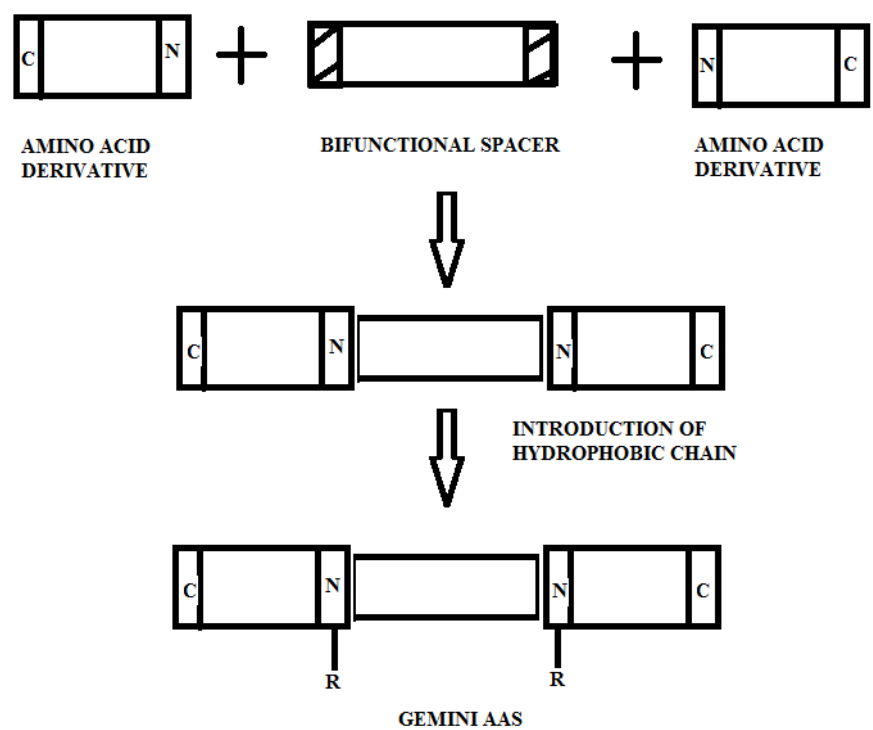
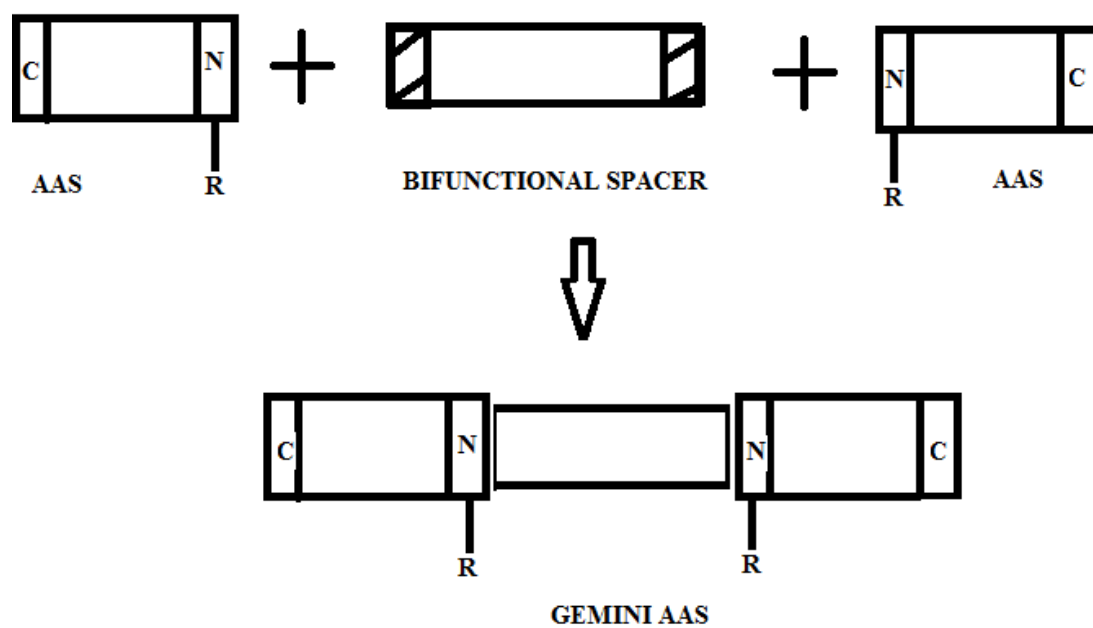


Figure 9: wt % of amino acid surfactants as the principal ingredients in shampoos, detergents and cosmetics

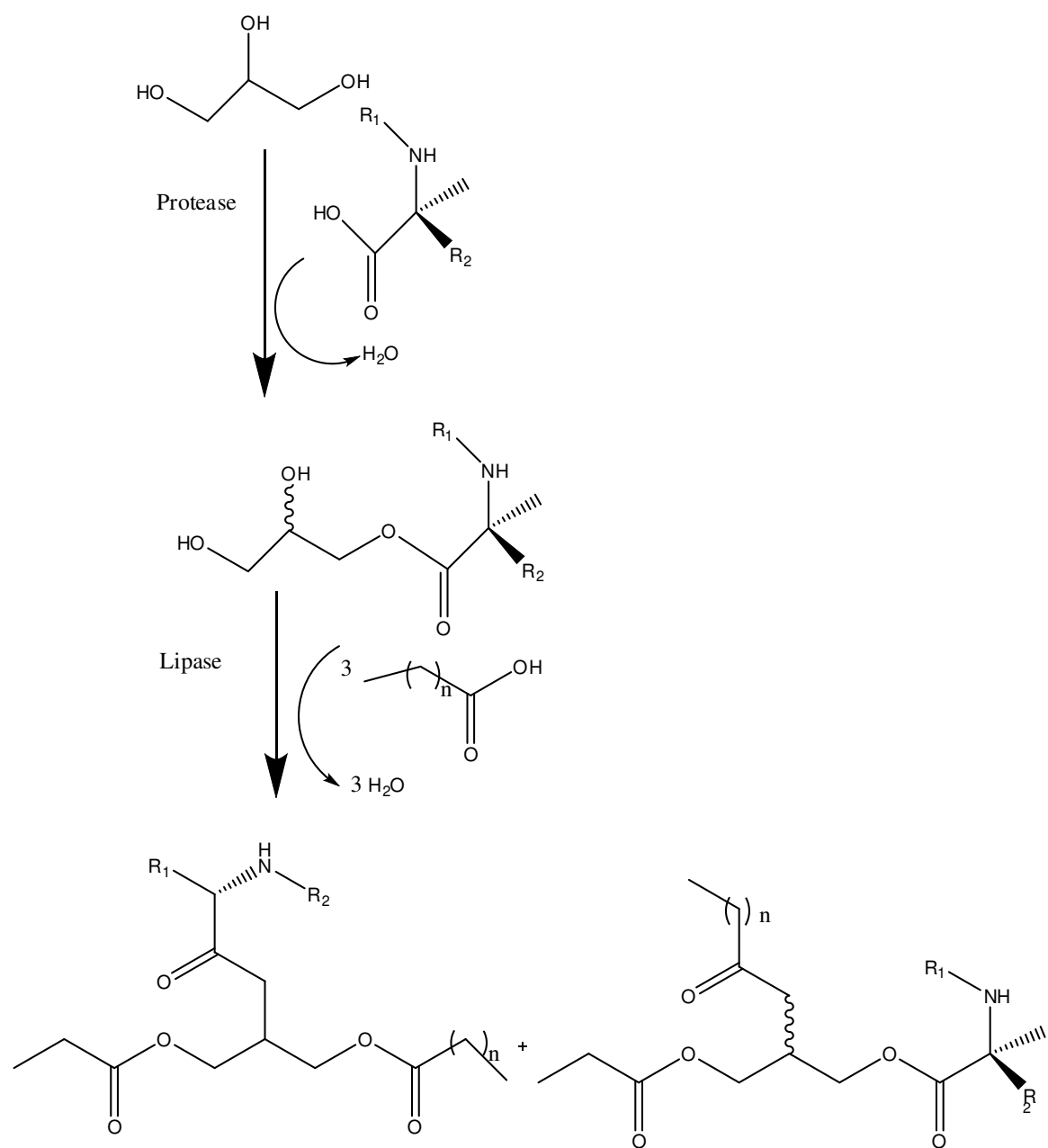
Figure



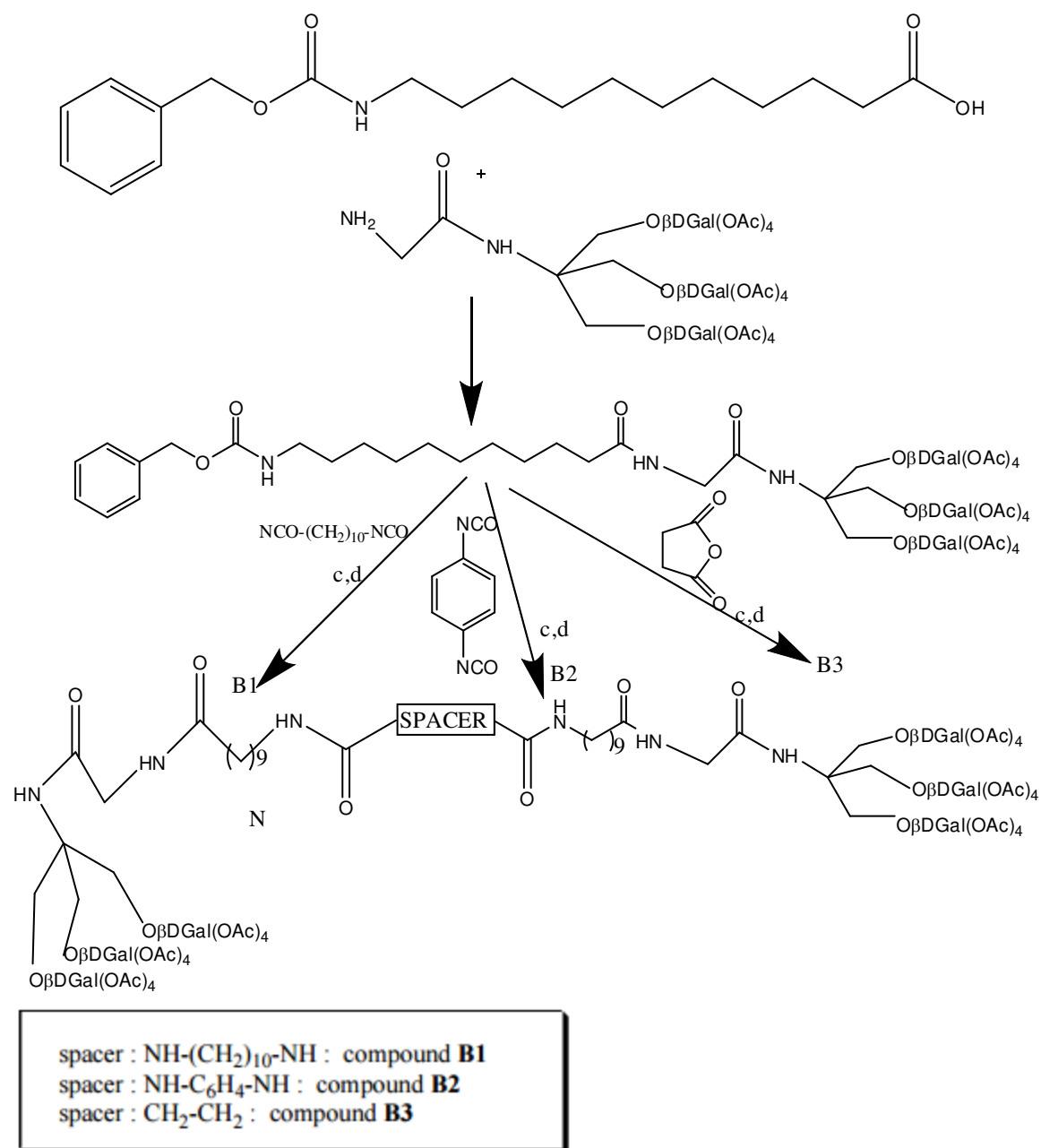
Scheme 1: Synthesis of Gemini AAS using AA derivatives and spacer, followed by insertion of hydrophobic group



Scheme 2: Synthesis of Gemini amino acid surfactants using bifunctional spacer and AAS



Scheme 3: Synthesis of mono and diacyl glycerol amino acid conjugates



Scheme 4: Synthesis of symmetric bolaamphiphiles derived from tris(hydroxymethyl)aminomethane (Tris)

	CAE	NAE	LAE	MAE
CMC (mg/L)	>1500	820±50	410±10	350±30
γ (mN/m)	27.0±0.5	26.1±0.5	25.5±0.5	24.0±0.5

Table 1: CMC and surface tension of N α -Acyl arginine surfactants

CAE= N- α -Octanoyl-L-Arginine ethyl ester, NAE= N- α -Nonanoyl-L-Arginine ethyl ester, LAE= Ethyl-N- α - lauroyl-L-arginate, MAE= N- α -Myristoyl-L-Arginine ethyl ester